AL/2020/02/E-I(NEW)

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NEW

අධ්‍යයන පොදු සහතික පත්‍ර (උසස් පෙළ) විභාගය, 2020
கல்விப் பொதுத் தராதரப் பத்திர (உயர் தர)ப் பரீட்சை, 2020
General Certificate of Education (Adv. Level) Examination, 2020

රසායන විද්‍යාව I இரசாயனவியல் I Chemistry I	02 E I	පැය දෙකයි இரண்டு மணித்தியாலம் Two hours

Instructions:

* *Periodic Table is provided.*
* *This paper consists of* **09** *pages.*
* *Answer* ***all*** *the questions.*
* ***Use of calculators is not allowed.***
* *Write your* ***Index Number*** *in the space provided in the answer sheet.*
* *Follow the instructions given on the back of the answer sheet carefully.*
* *In each of the questions* **1** *to* **50**, *pick one of the alternatives from* (1), (2), (3), (4), (5) *which is* ***correct or most appropriate*** *and* ***mark your response on the answer sheet with a cross*** (×) ***in accordance with the instructions given on the back of the answer sheet.***

Universal gas constant $R = 8.314\ J\ K^{-1}\ mol^{-1}$ Planck's constant $h = 6.626 \times 10^{-34}\ J\ s$

Avogadro constant $N_A = 6.022 \times 10^{23}\ mol^{-1}$ Velocity of light $c = 3 \times 10^{8}\ m\ s^{-1}$

1. Consider the following discoveries made with regard to the atomic structure.

I. Positive rays inside a cathode ray tube

II. Radioactivity by certain types of nuclei

The two scientists who discovered the above stated I and II respectively, are,

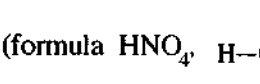
(1) J. J. Thomson and Henry Becquerel
(2) Eugen Goldstein and Robert Millikan
(3) Henry Becquerel and Eugen Goldstein
(4) J. J. Thomson and Ernest Rutherford
(5) Eugen Goldstein and Henry Becquerel

2. The number of electrons in the manganese atom (Mn, Z = 25) that have quantum numbers $l = 0$ and $m_l = -1$ respectively are,

(1) 6 and 4 (2) 8 and 12 (3) 8 and 5 (4) 8 and 6 (5) 10 and 5

3. **M** is an element that belongs to the second period in the Periodic Table. It forms a covalent molecule MCl_3 which has a dipole moment. The group of the Periodic Table to which **M** belongs is,

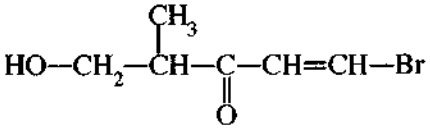
(1) 2 (2) 13 (3) 14 (4) 15 (5) 16

4. The number of **unstable** Lewis dot-dash structures that can be drawn for the peroxynitric acid molecule (formula HNO_4, ) is,

(1) 1 (2) 2 (3) 3 (4) 4 (5) 5

5. The IUPAC name of the given compound is,

(1) 1-bromo-4-methyl-5-hydroxypent-1-en-3-one
(2) 5-bromo-1-hydroxy-2-methylpent-4-en-3-one
(3) 1-bromo-5-hydroxy-4-methylpent-1-en-3-one
(4) 5-bromo-2-methyl-3-oxopent-4-en-1-ol
(5) 1-bromo-4-methyl-3-oxopent-1-enol



[See page two

6. The **decreasing** order of radii of the species O, O^{2-}, F, F^-, S^{2-}, Cl^- is,

(1) $S^{2-} > Cl^- > O^{2-} > F^- > O > F$
(2) $S^{2-} > Cl^- > O^{2-} > F^- > F > O$
(3) $Cl^- > S^{2-} > O^{2-} > F^- > O > F$
(4) $Cl^- > S^{2-} > F^- > O^{2-} > O > F$
(5) $S^{2-} > Cl^- > O^{2-} > O > F^- > F$

7. A rigid-closed container contains n_1 moles of an ideal gas at temperature T_1(K) and pressure P_1(Pa). When an additional amount of the gas was inserted into the container, the new temperature and pressure were T_2 and P_2, respectively. The total number of moles of the gas now in the container is,

(1) $\frac{n_1T_1P_1}{T_2P_2}$ (2) $\frac{n_1T_1P_2}{T_2P_1}$ (3) $\frac{T_2P_2}{n_1T_1P_1}$ (4) $\frac{n_1T_2P_2}{T_1P_1}$ (5) $\frac{n_1T_2P_1}{T_1P_2}$

8. The total number of electrons exchanged in the reaction of the oxidation of ethanol (C_2H_5OH) to acetic acid (CH_3COOH) using acidic $K_2Cr_2O_7$ solution is,

(1) 6 (2) 8 (3) 10 (4) 12 (5) 14

9. Which compound of the following, can undergo aldol condensation, when reacted with aqueous NaOH?

(1) $CH_3\overset{O}{\overset{\|}{C}}-OH$ (2) $CH_3\overset{O}{\overset{\|}{C}}-OCH_3$ (3) $H-\overset{O}{\overset{\|}{C}}-OCH_3$ (4) $CH_3CH_2\overset{O}{\overset{\|}{C}}-H$ (5) $(CH_3)_3C\overset{O}{\overset{\|}{C}}-H$

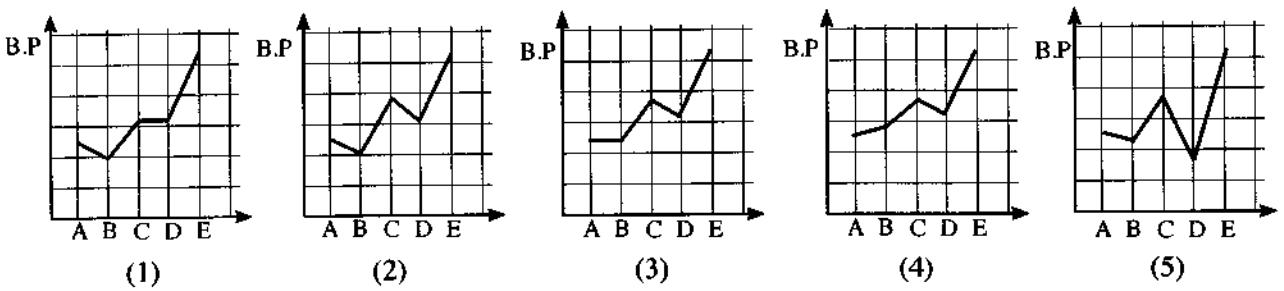
10. AX(s), A_2Y(s) and AZ(s) are sparingly soluble salts in water having K_{sp} values of 1.6×10^{-9}, 3.2×10^{-11} and 9.0×10^{-12}, respectively at 25 °C. Which of the following shows the order of the three saturated solutions of these salts in decreasing concentration of cation A^+(aq) , at 25 °C ?

(1) AX(s) > A_2Y(s) > AZ(s)
(2) A_2Y(s) > AX(s) > AZ(s)
(3) AX(s) > AZ(s) > A_2Y(s)
(4) A_2Y(s) > AZ(s) > AX(s)
(5) AZ(s) > A_2Y(s) > AX(s)

11. Consider the following compounds.

	A	B	C	D	E
	$CH_3CH_2CH_2CH_2CH_2CH_3$	$CH_3\overset{CH_3}{\overset{\|}{C}}CH_2CH_3$ (CH_3 below)	$CH_3CH_2CH_2CH_2CHO$	$CH_3\overset{CH_3}{\overset{\|}{C}}CHO$ (CH_3 below)	$CH_3CH_2CH_2CH_2CH_2OH$
Relative molecular mass	86	86	86	86	88

Variation of boiling points of these compounds is best shown by,



[See page three

12. The **increasing** order of covalent character of the chemical species NaCl, Na_2S, KF and KCl is,

(1) KF < NaCl < KCl < Na_2S
(2) KCl < NaCl < KF < Na_2S
(3) KF < KCl < NaCl < Na_2S
(4) Na_2S < NaCl < KCl < KF
(5) KF < Na_2S < NaCl < KCl

13. Standard combustion enthalpies of $H_2(g)$, C(s) and $CH_3OH(l)$ at 298 K are –286 kJ mol^{-1}, –393 kJ mol^{-1} and –726 kJ mol^{-1}, respectively. Enthalpy of vaporization of $CH_3OH(l)$ is +37 kJ mol^{-1}. Enthalpy of formation (kJ mol^{-1}) of one mole of **gaseous** CH_3OH at 298 K is,

(1) –276 (2) –239 (3) –202 (4) +84 (5) +202

14. Phosphorous can be prepared in an electric furnace as given by the following balanced chemical equation.

$$2\,Ca_3(PO_4)_2 + 6\,SiO_2 + 10\,C \longrightarrow 6\,CaSiO_3 + 10\,CO + P_4$$

When 620 g of $Ca_3(PO_4)_2$, 180 g of SiO_2 and 96 g of C were reacted, 50 g of P_4 were obtained. Under these conditions, the limiting reagent (reagent that is completely consumed) and percentage yield of P_4 respectively are, (C = 12, O = 16, Si = 28, P = 31, Ca = 40)

(1) $Ca_3(PO_4)_2$ and 80.7% (2) SiO_2 and 80.7% (3) C and 50.4%
(4) SiO_2 and 40.3% (5) C and 25.2%

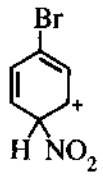
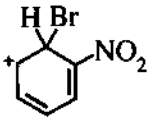
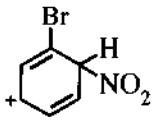
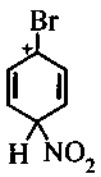
15. Consider the following two equilibria occurring in two separate rigid-closed containers under the same conditions.

$$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g) \;;\; K_{P_1} = 3.0 \times 10^{-4}$$
$$NH_3(g) + H_2S(g) \rightleftharpoons NH_4HS(g);\; K_{P_2} = 8.0 \times 10^{-4}$$

Under these conditions K_p for the equilibrium $2\,H_2S(g) + N_2(g) + 3\,H_2(g) \rightleftharpoons 2\,NH_4HS(g)$ is,

(1) 5.76×10^{-12} (2) 7.2×10^{-10} (3) 1.92×10^{-8} (4) 3.40×10^{-6} (5) 3.75×10^{-2}

16. Consider the nitration reaction of bromobenzene. Resonance stabilized carbocation intermediates are formed during this reaction. Which of the following is **not** a resonance structure of these intermediates?

(1) 
(2) 
(3) 
(4) 
(5) 

17. A reaction which is non-spontaneous at room temperature and 1 atm pressure becomes spontaneous at high temperature at the same pressure. Which of the following is correct for this reaction at room temperature? (Assume that ΔH and ΔS do not change with temperature and pressure.)

	ΔG	ΔH	ΔS
(1)	Positive	Positive	Positive
(2)	Positive	Negative	Negative
(3)	Positive	Negative	Positive
(4)	Negative	Positive	Negative
(5)	Negative	Negative	Negative

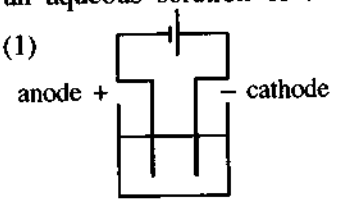
18. The de Broglie wavelength of a neutron travelling with a velocity v is λ. If the kinetic energy E ($E = \frac{1}{2}mv^2$) of this neutron is increased four times, the new de Broglie wavelength would be,

(1) $\frac{\lambda}{2}$ (2) $\frac{\lambda}{4}$ (3) 2λ (4) 4λ (5) 16λ

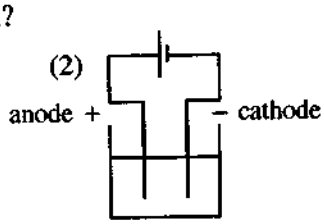
[See page four

19. Which of the following correctly shows the electrolytic cell constructed for the electrolysis of an aqueous solution of the salt MX?

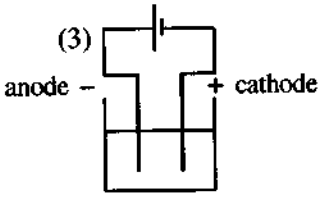
(1)



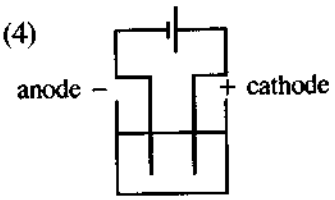
(2)



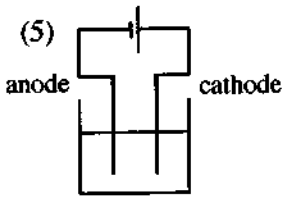
(3)



(4)



(5)



20. Which of the following statements is correct regarding the reaction between a carboxylic acid and an alcohol to give an ester?

(1) The overall reaction is a nucleophilic addition reaction of a carbonyl compound.

(2) It is a reaction in which the alcohol acts as a nucleophile.

(3) It is a reaction which occurs with the cleavage of the O–H bond of the carboxylic acid.

(4) It is a reaction which occurs with the cleavage of the C–O bond of the alcohol.

(5) It is an acid-base reaction.

21. Decomposition of 1 mol of $CH_3OH(l)$ occurs at high temperatures as follows.

$$CH_3OH(l) \rightarrow CO(g) + 2H_2(g);\ \Delta H = +128\,kJ$$

Which of the following is **incorrect** for the above reaction? (H = 1, C = 12, O = 16)

(1) The heat absorbed when 1 mol of $CH_3OH(g)$ is decomposed is less than 128 kJ.

(2) Enthalpy of $CO(g) + 2H_2(g)$ is higher than the enthalpy of $CH_3OH(l)$.

(3) 128 kJ of heat is released when 1 mol of CO(g) is formed.

(4) 128 kJ of heat is absorbed during the decomposition of a mole of reactant.

(5) 128 kJ of heat is absorbed when 32 g of products are formed.

22. Identify the **incorrect** statement from the following.

(1) Electron gain energy of nitrogen [N(g)] is positive.

(2) Dilution of $BiCl_3(aq)$ solution with water gives a white precipitate.

(3) H_2S gas can act both as an oxidizing agent and a reducing agent.

(4) The effective nuclear charge (Z*) felt by a valence electron in He is less than 2.

(5) Aluminium is inert towards N_2 gas even when heated to a high temperature.

23. The concentration of a dilute aqueous solution of a weak acid HA is C mol dm^{-3} and its acid dissociation constant is K_a at 298 K. Which of the following expressions gives the pH of the solution at 298 K?

(1) $pH = \frac{1}{2}pK_a - \frac{1}{2}\log C$

(2) $pH = -\frac{1}{2}pK_a - \frac{1}{2}\log C$

(3) $pH = -\frac{1}{2}pK_a + \frac{1}{2}\log C$

(4) $pH = -\frac{1}{2}pK_a - \frac{1}{2}\log(1/C)$

(5) $pH = \frac{1}{2}pK_a - \frac{1}{2}\log(1/C)$

[*See page five*

24. The strength of a H_2O_2 solution can be expressed as the volume of O_2 produced at standard temperature and pressure (STP). For example, a litre of 20 **volume strength** H_2O_2 solution will produce 20 litres of O_2 gas at STP ($2H_2O_2(aq) \rightarrow 2H_2O(l) + O_2(g)$). (Assume that 1 mole of gas has 22.4 litres volume at STP.)

A bottle labelled **X** contains H_2O_2 solution. When 25.0 cm^3 of solution **X** was titrated with 1.0 mol dm^{-3} $KMnO_4$ in the presence of dilute H_2SO_4 the volume required to reach the end point was 25.0 cm^3. The volume strength of solution **X** is,

(1) 15 (2) 20 (3) 25 (4) 28 (5) 30

25. $M(OH)_2(s)$ is a sparingly water soluble salt formed by the reaction between $M^{2+}(aq)$ and $OH^-(aq)$ ions at 298 K. The solubility (mol dm^{-3}) of $M(OH)_2(s)$ in water at pH = 5 is, ($K_{sp\,M(OH)_2} = 4.0 \times 10^{-36}$ at 298 K).

(1) $\sqrt{2} \times 10^{-18}$ (2) 2×10^{-18} (3) 1×10^{-18} (4) $\sqrt[3]{2} \times 10^{-12}$ (5) 1×10^{-12}

26. Which of the following correctly denotes the standard galvanic cell constructed by using a standard hydrogen electrode, a standard Mg–electrode and a salt-bridge at 298 K?

(1) $Mg(s) \mid Mg^{2+}(aq, 1.00 \text{ mol } dm^{-3}) \parallel H^+(aq, 1.00 \text{ mol } dm^{-3}) \mid H_2(g) \mid Pt(s)$

(2) $Pt(s) \mid H_2(g) \mid H^+(aq, 1.00 \text{ mol } dm^{-3}) \parallel Mg^{2+}(aq, 1.00 \text{ mol } dm^{-3}) \mid Mg(s)$

(3) $Mg(s), Mg^{2+}(aq, 1.00 \text{ mol } dm^{-3}) \parallel H^+(aq, 1.00 \text{ mol } dm^{-3}) \mid H_2(g) \mid Pt(s)$

(4) $Mg(s) \mid Mg^{2+}(aq, 1.00 \text{ mol } dm^{-3}), H^+(aq, 1.00 \text{ mol } dm^{-3}), H_2(g) \mid Pt(s)$

(5) $Pt(s), H_2(g) \mid H^+(aq, 1.00 \text{ mol } dm^{-3}) \parallel Mg^{2+}(aq, 1.00 \text{ mol } dm^{-3}), Mg(s)$

27. The following procedure was carried out at 298 K to determine the distribution coefficient K_D of a monobasic organic acid between dichloromethane and water. 50.00 cm^3 of a 0.20 mol dm^{-3} aqueous solution of acid were mixed vigorously with 10.00 cm^3 of dichloromethane and the two layers were allowed to separate. Thereafter, the dichloromethane layer in the bottom of the flask was drained out. 10.00 cm^3 of 0.02 mol dm^{-3} NaOH(aq) solution were required to neutralize the acid remaining in the aqueous layer. (Assume that the acid does not dimerize in the organic phase.) K_D of the acid between **dichloromethane and water** at 298 K is,

(1) 0.05 (2) 0.25 (3) 4.00 (4) 20.00 (5) 245.00

28. A reaction $C_2H_4(g) + 3O_2(g) \rightarrow 2CO_2(g) + 2H_2O(g)$ occurs in a rigid-closed container at a given temperature. After a certain time, it was found that the rate of the reaction with respect to consumption of $C_2H_4(g)$ was x mol $dm^{-3}\,s^{-1}$. Which of the following shows the rates of consumption of $O_2(g)$, formation of $CO_2(g)$ and formation of $H_2O(g)$ respectively, during that time?

	rate / mol $dm^{-3}\,s^{-1}$		
	$O_2(g)$	$CO_2(g)$	$H_2O(g)$
(1)	$\frac{3}{x}$	$\frac{2}{x}$	$\frac{2}{x}$
(2)	x	x	x
(3)	$\frac{x}{3}$	$\frac{x}{2}$	$\frac{x}{2}$
(4)	$\frac{1}{x}$	$\frac{1}{x}$	$\frac{1}{x}$
(5)	$3x$	$2x$	$2x$

29. Consider the following reaction occurring in a rigid-closed container at temperature T.

$$\mathbf{M}(g) + \mathbf{Q}(g) \rightarrow \mathbf{R}(g) + \mathbf{Z}(g)$$

The rate of reaction doubled when the concentration of **M** was doubled. The rate of reaction is 5.00×10^{-4} mol $dm^{-3}\,s^{-1}$ when the concentrations of **M** and **Q** are 1.0×10^{-5} mol dm^{-3} and 2.0 mol dm^{-3} respectively. The rate constant of the reaction under these conditions is,

(1) $2.5 \times 10^{-4}\,s^{-1}$ (2) $12.5\,s^{-1}$ (3) $25\,s^{-1}$ (4) $50\,s^{-1}$ (5) $500\,s^{-1}$

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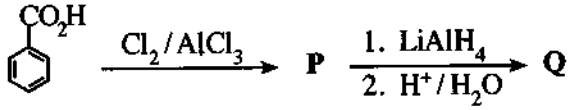
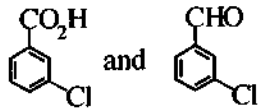
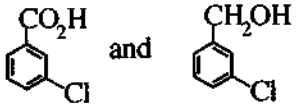
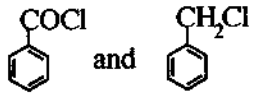
30. Consider the following reaction scheme.

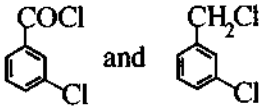


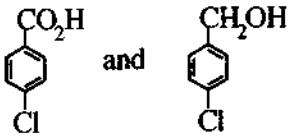
P and **Q** respectively could be,

(1) 

(2) 

(3) 

(4) 

(5) 

● For each of the questions **31** to **40**, one or more responses out of the four responses (*a*), (*b*), (*c*) and (*d*) given is/are correct. Select the correct response/responses. In accordance with the instructions given on your answer sheet, mark

(1) if only (*a*) and (*b*) are correct.
(2) if only (*b*) and (*c*) are correct.
(3) if only (*c*) and (*d*) are correct.
(4) if only (*d*) and (*a*) are correct.
(5) if **any other** number or combination of responses is correct.

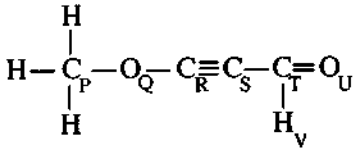
Summary of above Instructions

(1)	(2)	(3)	(4)	(5)
Only (*a*) and (*b*) are correct	Only (*b*) and (*c*) are correct	Only (*c*) and (*d*) are correct	Only (*d*) and (*a*) are correct	**Any other** number or combination of responses is correct

31. Which of the following statement/s is/are correct with regard to 3*d*-block elements and their compounds?

(*a*) Among the 3*d*-block elements, Sc is not considered as a transition element.
(*b*) The radii of atoms (Sc to Cu) decrease from left to right.
(*c*) $[Ni(NH_3)_6]^{2+}$ is blue in colour whereas $[Zn(NH_3)_4]^{2+}$ is colourless.
(*d*) The IUPAC name of K_2NiCl_4 is dipotassium tetrachloronickelate(II).

32. Which statement/s is/are correct regarding the following molecule?



(*a*) Atoms labelled P, Q, R and S lie on a straight line.
(*b*) Atoms labelled Q, R, S and T lie on a straight line.
(*c*) Atoms labelled R, S, T, U and V lie on the same plane.
(*d*) Atoms labelled R, S, T and U lie on a straight line.

33. 0.01 moles of $N_2(g)$, 0.10 moles of $H_2(g)$ and 0.40 moles of $NH_3(g)$ were inserted into a 1.0 dm^3 rigid-closed container and allowed to reach equilibrium at 500 K as given below.

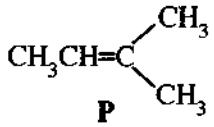
$$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g) \qquad K_C = 2.0 \times 10^2\ mol^{-2}\ dm^6$$

Which of the following statement/s is/are correct for the changes in the system from the initial stage to equilibrium? Q_C is the reaction quotient.

(*a*) Initially $Q_C > K_C$; $NH_3(g)$ starts to produce $N_2(g)$ and $H_2(g)$ and the system reaches equilibrium.
(*b*) Initially $Q_C < K_C$; $NH_3(g)$ starts to produce $N_2(g)$ and $H_2(g)$ and the system reaches equilibrium.
(*c*) Initially $Q_C < K_C$; $N_2(g)$ and $H_2(g)$ react to form $NH_3(g)$ and the system reaches equilibrium.
(*d*) Initially $Q_C > K_C$; $N_2(g)$ and $H_2(g)$ react to form $NH_3(g)$ and the system reaches equilibrium.

[*See page seven*

34. Which of the following statement/s regarding the reaction between compound **P** and HCl to form an alkyl halide is/are correct?



P

(*a*) The major product is 2-chloro-2-methylbutane.
(*b*) A secondary carbocation is formed as an intermediate in this reaction.
(*c*) In one of the steps of the reaction, the HCl bond is cleaved to give a chlorine radical ($Cl^{\bullet}$).
(*d*) In one of the steps of the reaction, a nucleophile reacts with a carbocation.

35. A binary liquid mixture prepared by mixing two liquids in a closed evacuated container at a given temperature shows a negative deviation from Roult's Law. Which of the following statement/s is/are correct for this system?
(*a*) Total vapour pressure of the mixture is less than the expected total vapour pressure should it behave as an ideal mixture.
(*b*) Heat is released when the mixture is formed.
(*c*) Number of molecules in the vapour phase of the mixture is greater than the expected number of molecules should it behave as an ideal mixture.
(*d*) Heat is absorbed when the mixture is formed.

36. Which of the following statement/s is/are correct with regard to CFC, HCFC and HFC?
(*a*) Both classes of compounds CFC and HCFC have the ability to produce chlorine free radicals in the upper atmosphere (stratosphere).
(*b*) Both classes of compounds HFC and HCFC have the ability to produce chlorine free radicals in the upper atmosphere (stratosphere).
(*c*) All three classes of compounds CFC, HCFC and HFC are strong greenhouse gases.
(*d*) All three classes of compounds CFC, HCFC and HFC contribute significantly to ozone layer depletion.

37. Which of the following statement/s is/are correct with regard to halogens, noble gases and their compounds?
(*a*) Hypochlorous ion disproportionates rapidly in acidic solutions.
(*b*) Xe forms a series of compounds with F_2 gas, among which XeF_4 has a square planar geometry.
(*c*) Among the hydrogen halides, HF has the highest bond dissociation energy per mole.
(*d*) Boiling points of halogens increase down the group as a result of increasing strength of London forces.

38. Which of the following statement/s is/are correct regarding the Daniell cell when it operates at room temperature? (E°_{cell} = +1.10 V)
(*a*) Net electron flow occurs from Zn to Cu.
(*b*) The equilibrium $Zn^{2+}(aq) + 2e \rightleftharpoons Zn(s)$ shifts to the right.
(*c*) A liquid-junction potential is created due to the presence of a salt-bridge.
(*d*) The equilibrium $Cu^{2+}(aq) + 2e \rightleftharpoons Cu(s)$ shifts to the right.

39. Which of the following statement/s is/are correct for ideal gases and real gases at constant temperature?
(*a*) At very high pressures, the volume of a real gas is higher than that of an ideal gas.
(*b*) At high pressures, real gases tend to behave as ideal gases.
(*c*) At very high pressures, the volume of a real gas is lower than that of an ideal gas.
(*d*) At low pressures, real gases tend to behave as ideal gases.

40. Which of the following statement/s is/are correct regarding some industrial processes?
(*a*) The first two steps involved in the manufacture of Na_2CO_3 by Solvay Process are endothermic.
(*b*) The presence of Mg^{2+}, Ca^{2+} and SO_4^{2-} ions in brine, hinders the production of NaOH using the membrane cell method.
(*c*) The first step involved in the manufacture of nitric acid by Ostwald method is the oxidation of NH_3 gas using O_2 in air in the presence of a catalyst to give NO_2 gas.
(*d*) High temperature and low pressure conditions are employed in the manufacture of NH_3 gas using Haber-Bosh process.

[*See page eight*

- In question Nos. **41** to **50**, two statements are given in respect of each question. From the Table given below, select the response, out of the responses (1), (2), (3), (4) and (5), that **best** fits the two statements and mark appropriately on your answer sheet.

Response	First Statement	Second Statement
(1)	True	True, and correctly explains the first statement
(2)	True	True, but does **not** explain the first statement correctly
(3)	True	False
(4)	False	True
(5)	False	False

	First Statement	Second statement
41.	Among the oxides of Cr and Mn, CrO and MnO are acidic, while CrO_3 and Mn_2O_7 are basic.	The acidic/basic nature of the oxides of Cr and Mn is dependant on the oxidation number of the metal.
42.	An acidic buffer solution can be prepared by mixing a weak acid HA(aq) with its sodium salt NaA(aq).	When $OH^-(aq)$ or $H^+(aq)$ ions are added to a buffer solution, the added amounts of $OH^-(aq)$ or $H^+(aq)$ ions are removed through the reactions; $OH^-(aq) + HA(aq) \rightarrow A^-(aq) + H_2O(l)$ and $H^+(aq) + A^-(aq) \rightarrow HA(aq)$ respectively.
43.	Essential oils can be extracted from plants by steam distillation at a temperature below 100 °C.	At the temperature at which a mixture of essential oil and water boils, the total vapour pressure of the system is less than the atmospheric pressure.
44.	At a given temperature and pressure the molar volumes of two different ideal gases are different from each other.	At 0 °C temperature and 1 atm pressure, the molar volume of an ideal gas is 22.4 $dm^3\ mol^{-1}$.
45.	All compounds having a C=C bond show diastereoisomerism.	Any two isomers which are not mirror images of each other are diastereoisomers.
46.	Hydrogenation of benzene is more difficult than hydrogenation of alkenes.	Addition of hydrogen to benzene results in the loss of aromatic stabilization.
47.	The reaction that takes place between SO_3 gas and water in the production of sulphuric acid is endothermic.	SO_3 gas reacts with concentrated H_2SO_4 to give oleum.
48.	Reaction between ammonia and an alkylhalide gives a mixture of primary, secondary and tertiary amines and a quaternary ammonium salt.	Primary, secondary and tertiary amines can react as nucleophiles.
49.	If $P + Q \rightarrow R$ is a first order reaction with respect to the reactant P, the graph of rate against concentration of P gives a straight line passing through the origin.	Initial rate of a first order reaction is independent of the concentration of reactant(s).
50.	On a sunny day, strong photochemical smog can be seen in a city with heavy traffic congestion.	Photochemical smog is caused entirely by scattering of solar radiation by small particles and water droplets that are emitted by vehicle exhaust systems.

* * *

[*See page nine*

නව නිර්දේශය/புதிய பாடத்திட்டம்/New Syllabus

NEW

අධ්‍යයන පොදු සහතික පත්‍ර (උසස් පෙළ) විභාගය, 2020
கல்விப் பொதுத் தராதரப் பத்திர (உயர் தர)ப் பரீட்சை, 2020
General Certificate of Education (Adv. Level) Examination, 2020

රසායන විද්‍යාව II இரசாயனவியல் II **Chemistry II**	**02 E II**

පැය තුනයි மூன்று மணித்தியாலம் ***Three hours***	අමතර කියවීම් කාලය - මිනිත්තු 10 යි மேலதிக வாசிப்பு நேரம் - 10 நிமிடங்கள் **Additional Reading Time - 10 minutes**

Use **additional reading time** to go through the question paper, select the questions and decide on the questions that you give priority in answering.

Index No. : .....

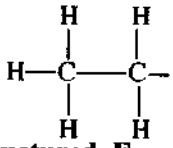
* *A Periodic Table is provided on page 15.*

* ***Use of calculators is not allowed.***

* *Universal gas constant,* $R = 8.314\ J\ K^{-1}\ mol^{-1}$

* *Avogadro constant,* $N_A = 6.022 \times 10^{23}\ mol^{-1}$

* **In answering this paper, you may represent alkyl groups in a condensed manner.**

Example:  group may be shown as CH_3CH_2-

❑ **PART A — Structured Essay (pages 02 - 08)**

* *Answer **all** the questions on the question paper itself.*

* *Write your answer in the space provided for each question. Please note that the space provided is sufficient for the answer and that extensive answers are not expected.*

❑ **PART B and PART C — Essay (pages 09 - 14)**

* *Answer **four** questions selecting **two** questions from each part. Use the papers supplied for this purpose.*

* *At the end of the time allotted for this paper, tie the answers to the three Parts **A**, **B** and **C** together so that Part **A** is on top and hand them over to the Supervisor.*

* *You are permitted to remove **only** Parts **B** and **C** of the question paper from the Examination Hall.*

For Examiner's Use Only

Part	Question No.	Marks
A	1	
	2	
	3	
	4	
B	5	
	6	
	7	
C	8	
	9	
	10	
Total		

Total

In Numbers	
In Letters	

Code Numbers

Marking Examiner 1	
Marking Examiner 2	
Checked by :	
Supervised by :	

[see page two

PART A — STRUCTURED ESSAY

*Answer **all four** questions on this paper itself. (Each question carries **100** marks.)*

Do not write in this column.

1. (*a*) Write the answers to the questions given below on the dotted lines.

(i) Of the three ions Na^+, Mg^{2+} and F^-, which one has the **smallest** ionic radius? .....

(ii) Of the three elements C, N and O, which one has the **highest** second ionization energy? .....

(iii) Of the three compounds H_2O, HOCl and OF_2, which one has the **most** electronegative oxygen atom? .....

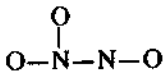
(iv) Of the three elements Be, C and N, which one will liberate energy when an electron is added to its atom [$Y(g) + e \rightarrow Y^-(g)$; Y = Be, C, N] in the gaseous state? .....

(v) Of the three ionic compounds NaF, KF and KBr, which one has the **highest** solubility in water? .....

(vi) Of the three compounds HCHO, CH_3F and H_2O_2, which one has the **strongest** intermolecular forces? .....

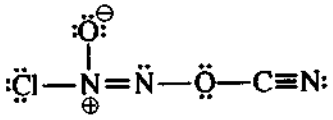
(24 *marks*)

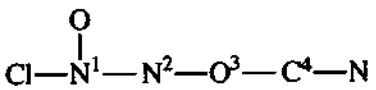
(*b*) (i) Draw the **most** acceptable Lewis dot-dash structure for the ion, $N_2O_3^{2-}$. Its skeleton is given below.



(ii) Draw **three** more Lewis dot-dash structures (resonance structures) for this ion. Indicate the relative stabilities of the structures drawn by you, when compared with the most acceptable structure drawn in (i) above, by writing **'less stable'** or **'unstable'** under these structures.

(iii) Complete the given table based on the Lewis dot-dash structure and its labelled skeleton given below.





	N^1	N^2	O^3	C^4
VSEPR pairs around the atom				
electron pair geometry around the atom				
shape around the atom				
hybridization of the atom				

[*see page three*

Do not write in this column.

- Parts (iv) to (vii) are based on the Lewis dot-dash structure given in part (iii) above. Labelling of atoms is as in part (iii).

(iv) Identify the atomic / hybrid orbitals involved in the formation of σ bonds between the two atoms given below.

I.	$Cl—N^1$	Cl	N^1	.....
II.	$N^1—O$	N^1	O	.....
III.	$N^1—N^2$	N^1	N^2	.....
IV.	$N^2—O^3$	N^2	O^3	.....
V.	$O^3—C^4$	O^3	C^4	.....
VI.	$C^4—N$	C^4	N	.....

(v) Identify the atomic orbitals involved in the formation of π bonds between the two atoms given below.

I.	$N^1—N^2$	N^1	N^2	.....
II.	$C^4—N$	C^4	N	.....
		C^4	N	.....

(vi) State the approximate bond angles around N^1, N^2, O^3 and C^4 atoms.

N^1....., N^2 ....., O^3 ....., C^4.....

(vii) Arrange the atoms N^1, N^2, O^3 and C^4 in the **increasing** order of electronegativity.

..... < ..... < ..... < ..... ***(56 marks)***

(c) Consider the following information.

I. The atoms **A** and **B** combine to form a heterodiatomic molecule **AB** that has a σ bond. This is represented as **A–B**.

II. The electronegativity of **A** is less than that of **B** ($X_A < X_B$).
X = electronegativity of the atom

III. The inter-nuclear distance between **A** and **B** atoms (d_{A-B}) of the **AB** molecule is given by the following equation.

$$d_{A-B} = r_A + r_B - c(X_B - X_A)$$

r = atomic radius, c = 9 pm

Note: d and r are measured in picometres (pm). (1 pm = 10^{-12} m)

Based on the above information, answer the following questions.

(i) What is the name used to identify the type of σ bond between **A** and **B**?

.....

(ii) Show how fractional charges (δ+ and δ–) are located in the molecule **AB**.

.....

(iii) Write the equation to calculate the dipole moment (μ) of molecule **AB** and show its direction.

[see page four

Do not write in this column.

(iv) Calculate the percentage of ionic character of the H–F bond in the HF molecule using the data given below.

Inter-nuclear distance of H_2 (d_{H-H}) = 74 pm
Inter-nuclear distance of F_2 (d_{F-F}) = 144 pm
Electronegativity of H = 2.1
Electronegativity of F = 4.0
Dipole moment of HF = 6.0×10^{-30} C m
Charge of an electron = 1.6×10^{-19} C

100

(20 *marks*)

2. (*a*) **A, B, C** and **D** are chlorides of *p*-block elements. These elements have atomic numbers less than 20. A description of the products ($\mathbf{P_1} - \mathbf{P_9}$) formed when **A** is reacted with a limited amount of water and **B, C** and **D** are reacted with excess water are given below.

Compound		Description of products
A	$\mathbf{P_1}$	a compound with a covalent network structure
	$\mathbf{P_2}$	a strong monobasic acid
B	$\mathbf{P_3}$	a gas that turns red litmus blue
	$\mathbf{P_4}$	a compound with bleaching properties
C	$\mathbf{P_5}$	a tribasic acid
	$\mathbf{P_6}$	a strong monobasic acid
D	$\mathbf{P_7}$	a gas that turns acidic $KMnO_4$ solution colourless
	$\mathbf{P_8}$	a colloidal solid
	$\mathbf{P_9}$	a strong monobasic acid

(i) Identify **A, B, C** and **D** (give the chemical formulae).

A: ..... **B:** ..... **C:** ..... **D:** .....

(ii) Give balanced chemical equations for the reactions of **A, B, C** and **D** with water to give products $\mathbf{P_1}$ to $\mathbf{P_9}$.

.....

.....

.....

.....

[*see page five*

Do not write in this column.

(iii) Write balanced chemical equations for the following reactions.

I. $\mathbf{P_1}$ with NaOH(aq)

.....

II. $\mathbf{P_3}$ with Mg

.....

III. $\mathbf{P_7}$ with acidic $K_2Cr_2O_7$

.....

(50 *marks*)

(*b*) A student is provided with bottles labelled **P, Q, R, S, T** and **U** containing aqueous solutions of $Al_2(SO_4)_3$, H_2SO_4, $Na_2S_2O_3$, $BaCl_2$, $Pb(Ac)_2$ and KOH **(not in order)**. Some useful observations for their identification on mixing two solutions at a time are given below. (Ac - Acetate ion)

	Solutions mixed	Observations
I	T + R	a clear colourless solution
II	P + R	a white precipitate
III	T + S	a gelatinous white precipitate
IV	U + R	a white precipitate
V	P + Q	a white precipitate, turns black on heating
VI	P + U	a white precipitate, dissolves on heating

(i) Identify **P** to **U**.

P: ..... **Q**: ..... **R**: .....

S: ..... **T**: ..... **U**: .....

(ii) Give balanced chemical equations for each of the reactions **I** to **VI**.

I: .....

II: .....

III: .....

IV: .....

V: formation of white precipitate: .....

turning black on heating: .....

VI: .....

(**Note**: indicate precipitates as ↓)

(50 *marks*)

100

3. (*a*) A saturated aqueous solution of a sparingly soluble salt $AB_2(s)$ was prepared by stirring an excess amount of $AB_2(s)$ in 1.0 dm^3 of distilled water at 25 °C. The amount of $A^{2+}(aq)$ ions present in this saturated aqueous solution was found to be 2.0×10^{-3} mol.

(i) Write the equilibrium related to the dissolution of $AB_2(s)$ in the above system at 25 °C.

.....

(ii) Write the expression for the equilibrium constant for the equilibrium written in (i) above at 25 °C.

.....

.....

[*see page six*

Do not write in this column.

(iii) Calculate the value of the equilibrium constant stated in (ii) above at 25 °C.

.....

.....

.....

.....

.....

.....

(iv) Another saturated aqueous solution of AB_2 was prepared by stirring an excess amount of $AB_2(s)$ in 2.0 dm^3 of distilled water at 25 °C. Giving reasons, predict the value of the equilibrium constant for this system.

.....

.....

(v) A small amount of the strong electrolyte NaB(s) is added to a saturated aqueous solution of AB_2 at 25 °C. Giving reasons, predict whether the concentration of $A^{2+}(aq)$ is increased or decreased.

.....

.....

.....

(60 *marks*)

(*b*) In an aqueous solution, propanoic acid (C_2H_5COOH) ionizes as given below.

$$C_2H_5COOH(aq) + H_2O(l) \rightleftharpoons C_2H_5COO^-(aq) + H_3O^+(aq)$$

At 25 °C, K_a (propanoic acid) = 1.0×10^{-5}

(i) Write the expression for the equilibrium constant for the above reaction at 25 °C.

.....

.....

(ii) 100.0 cm^3 of an aqueous solution of $C_2H_5COOH(aq)$ was prepared by dissolving 0.74 cm^3 of C_2H_5COOH in distilled water at 25 °C. Calculate the pH of the solution at 25 °C. (C = 12; O = 16; H = 1; consider the density of C_2H_5COOH as 1.0 g cm^{-3})

(40 *marks*)

100

[*see page seven*

Do not write in this column.

4. (*a*) **A, B, C** and **D** are structural isomers having the molecular formula C_6H_{10}. None of them show optical isomerism. All four isomers, **A, B, C** and **D** when treated with $HgSO_4$/dil. H_2SO_4 give products which react with 2,4-dinitrophenylhydrazine (2,4-DNP) to give coloured precipitates.

Only **A** gives a precipitate with ammonical $AgNO_3$. **A** has only one position isomer, which is **B**. **B** is a chain isomer of **C**. **C** reacts with $HgSO_4$/dil. H_2SO_4 to give two products **E** and **F**. **D** reacts with $HgSO_4$/dil. H_2SO_4 to give only one product, which is **E**.

(i) Draw the structures of **A, B, C, D, E** and **F** in the boxes given below.

A **B** **C**

D **E** **F**

(ii) Which of the compounds **A, B, C** and **D** gives a product that does not show diastereoisomerism when reacted separately with H_2 / Pd-$BaSO_4$ / quinoline?

(iii) Draw, in the box given below, the structure of the product **G** obtained when **A** is reacted with excess HBr.

G

(iv) Draw the structures of products **X** and **Y** obtained in the following reactions of **E**, in the appropriate boxes.

$$\text{X} \xleftarrow[\text{Methanol}]{NaBH_4} \text{E} \xrightarrow[2.\ H^+/H_2O]{1.\ C_2H_5MgBr} \text{Y}$$

Name a test to distinguish between **X** and **Y**.

.....

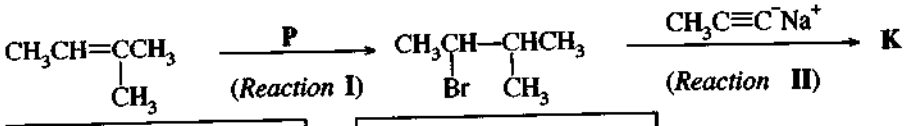
(60 *marks*)

[*see page eight*

Do not write in this column.

(*b*) (i) Complete the following three reaction sequences by drawing structures of compounds **K, L** and **M** and giving the reagents/catalysts **P, Q** and **R** in the boxes given below.

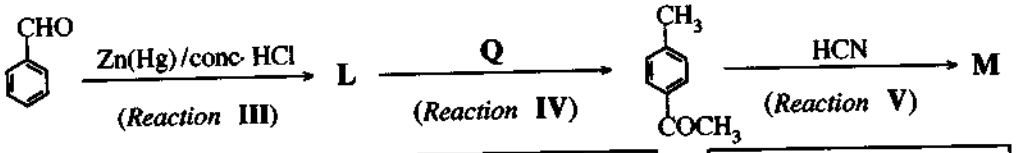
Sequence 1:



P

K

Sequence 2:

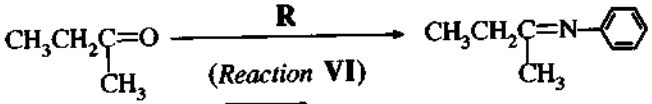


L

Q

M

Sequence 3:



R

(30 *marks*)

(ii) Selecting from the reactions **I – VI**, give **one (01)** example for each of the following types of reactions.

Nucleophilic addition .....

Nucleophilic substitution ..... **(10 *marks*)**

100

* *

[*see page nine*

නව නිර්දේශය/புதிய பாடத்திட்டம்/New Syllabus

NEW

අධ්‍යයන පොදු සහතික පත්‍ර (උසස් පෙළ) විභාගය, 2020
கல்விப் பொதுத் தராதரப் பத்திர (உயர் தர)ப் பரீட்சை, 2020
General Certificate of Education (Adv. Level) Examination, 2020

රසායන විද්‍යාව	II
இரசாயனவியல்	II
Chemistry	**II**

02 E II

* Universal gas constant $R = 8.314\ J\ K^{-1}\ mol^{-1}$

* Avogadro constant $N_A = 6.022 \times 10^{23}\ mol^{-1}$

PART B — ESSAY

*Answer **two** questions only. (Each question carries **150 marks**.)*

5. (*a*) A compound $XY_2Z_2(g)$ undergoes dissociation when heated to temperatures above 300 K as given below.

$$XY_2Z_2(g) \overset{\Delta}{\rightleftharpoons} XY_2(g) + Z_2(g)$$

A sample of 7.5 g of $XY_2Z_2(g)$ was placed in an evacuated 1.00 dm^3 rigid-closed container and the temperature was raised to 480 K.

Molar mass of $XY_2Z_2(g)$ is 150 g mol^{-1}. Use the approximate value of 4000 J mol^{-1} for RT at 480 K. Assume ideal gas behaviour for all gases.

(i) Calculate the number of moles of $XY_2Z_2(g)$ in the container before dissociation.

(ii) When the above system reaches equilibrium at 480 K, the total number of moles in the container was found to be 7.5×10^{-2} mol. Calculate the number of moles of $XY_2Z_2(g)$, $XY_2(g)$ and $Z_2(g)$ in the equilibrium mixture at 480 K.

(iii) Calculate the equilibrium constant K_c for the above reaction at 480 K.

(iv) Calculate K_p for the equilibrium at 480 K. **(75 *marks*)**

(*b*) For the reaction $XY_2Z_2(g) \rightarrow XY_2(g) + Z_2(g)$ described in (a), Gibbs free energies (G) at 480 K for $XY_2Z_2(g)$, $XY_2(g)$ and $Z_2(g)$ are –60 kJ mol^{-1}, –76 kJ mol^{-1} and –30 kJ mol^{-1}, respectively.

(i) Calculate ΔG (in kJ mol^{-1}) for the reaction at 480 K.

(ii) The magnitude of ΔS of the above reaction is 150 J K^{-1} mol^{-1} at 480 K. Calculate ΔH for the reaction at 480 K by using the appropriate sign (– or +) of ΔS.

(iii) By using the sign (– or +) of ΔH obtained in (ii), explain whether this reaction is exothermic or endothermic.

(iv) Deduce the enthalpy difference for the formation of $XY_2Z_2(g)$ from $XY_2(g)$ and $Z_2(g)$ at 480 K.

(v) If the bond enthalpy of the X–Z bond in $XY_2Z_2(g)$ is +250 kJ mol^{-1}, calculate the bond enthalpy of the Z–Z bond.

(Assume that $XY_2Z_2(g)$ has the structure $Z-\overset{\overset{Y}{\|}}{\underset{\underset{Y}{\|}}{X}}-Z$)

(vi) If liquid XY_2Z_2 is used instead of gaseous XY_2Z_2, giving reasons, explain whether the value of ΔH obtained for the reaction $XY_2Z_2(l) \rightarrow XY_2(g) + Z_2(g)$ is equal to, or higher or lower than ΔH obtained in (ii).

(75 *marks*)

[*see page ten*

6. (*a*) Consider the reaction given below occurring in a closed container at a given temperature T.

$$2N_2O_5(g) \rightarrow 4NO_2(g) + O_2(g)$$

(i) Write **three** expressions for the rate of reaction relevant to each of the compounds appearing in the reaction.

(ii) This reaction was carried out at temperature T with an initial concentration of 0.10 mol dm^{-3} of $N_2O_5(g)$. It was found that 40% of the initial amount was decomposed after a period of 400 s.

I. Calculate the average rate of decomposition of $N_2O_5(g)$ in this time interval.

II. Calculate average rates of formation of $NO_2(g)$ and $O_2(g)$.

(iii) In another experiment, initial rates were measured for this reaction at 300 K and the results are given below.

$[N_2O_5(g)]$ / mol dm^{-3}	0.01	0.02	0.03
Initial rate / mol dm^{-3} s^{-1}	6.930×10^{-5}	1.386×10^{-4}	2.079×10^{-4}

Derive the rate law for the reaction at 300 K.

(iv) Another experiment was carried out at 300 K with an initial concentration of 0.64 mol dm^{-3} of $N_2O_5(g)$. It was found that the concentration of $N_2O_5(g)$ which remained after a period of 500 s was 2.0×10^{-2} mol dm^{-3}.

I. Calculate the half-life ($t_{1/2}$) of the reaction at 300 K.

II. Calculate the rate constant of the reaction at 300 K.

(v) This reaction proceeds through a mechanism involving the following elementary steps.

Step 1 : $N_2O_5(g) \rightleftharpoons NO_3(g) + NO_2(g)$: Fast

Step 2 : $NO_3(g) + NO_2(g) \rightarrow 2NO_2(g) + O(g)$: Slow

Step 3 : $N_2O_5(g) + O(g) \rightarrow 2NO_2(g) + O_2(g)$: Fast

Show that the above mechanism is consistent with the rate law of the reaction. **(80 *marks*)**

(*b*) An ideal binary-liquid mixture was prepared by mixing two liquids of **A** and **B** in a closed evacuated container at temperature T. After establishing the equilibrium at temperature T, partial pressures of **A** and **B** in the vapour phase are P_A and P_B, respectively. At temperature T, the saturated vapour pressures of **A** and **B** are P_A° and P_B°, respectively. Mole fractions of **A** and **B** in solution are X_A and X_B, respectively.

(i) Show that $P_A = P_A^\circ X_A$

(Consider that the rates of vaporization and condensation are equal at equilibrium.)

(ii) In the above system at 300 K, the total pressure was 5.0×10^4 Pa. The saturated vapour pressures of pure **A** and **B** at 300 K, are 7.0×10^4 Pa and 3.0×10^4 Pa, respectively.

I. Calculate the mole fraction of **A** in the liquid phase of the equilibrium mixture.

II. Calculate the vapour pressure of **A** in the equilibrium mixture.

(70 *marks*)

[*see page eleven*

7. (*a*) (i) To compare the properties of Electrolytic and Galvanic cells, copy and complete the following table using the given terms.

Terms: anode, cathode, positive, negative, spontaneous, non-spontaneous.

		Electrolytic cell	Galvanic cell
A.	Oxidation half-reaction takes place at		
B.	Reduction half-reaction takes place at		
C.	Sign of E°_{cell}		
D.	Electron flow	From ..... to	From ..... to .....
E.	Spontaneity of the cell reaction		

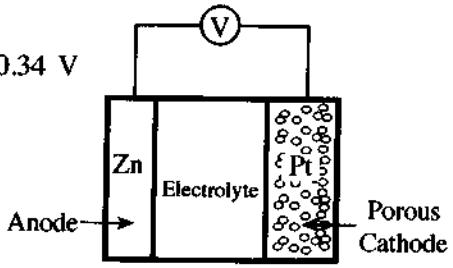
(ii) An electrochemical cell was constructed at 300 K by using a Zn(s) anode, an aqueous alkaline electrolyte and a porous Pt cathode which facilitates the collection of oxygen $O_2(g)$ from air as shown below. As the cell operates ZnO(s) is produced.

You are given that

$E^{\circ}_{ZnO(s)\,|\,Zn(s)\,|\,OH^-(aq)} = -1.31$ V and $E^{\circ}_{O_2(g)\,|\,OH^-(aq)} = +0.34$ V

Zn = 65 g mol^{-1}, O = 16 g mol^{-1} and

1 *F* = 96,500 C



I. Write the half-reactions occurring at anode and cathode.

II. Write the overall cell reaction.

III. Calculate the cell potential E°_{cell} at 300 K.

IV. State the direction of migration of $OH^-(aq)$ ions between the electrodes.

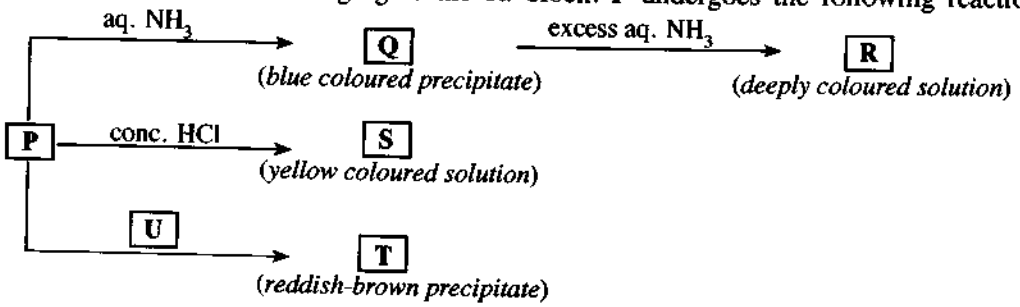
V. When the cell operates for a period of 800 s at 300 K, 2 mol of $O_2(g)$ are consumed.

A. Calculate the number of moles of electrons passing through the cell.

B. Calculate the mass of ZnO(s) formed.

C. Calculate the current passing through the cell. **(75 *marks*)**

(*b*) A coloured complex ion **P** is formed when the salt $M(NO_3)_n$ is dissolved in distilled water. **M** is a transition element belonging to the 3*d* block. **P** undergoes the following reactions.



T and **U** are coordination compounds each containing four elements. **P, R** and **S** are complex ions.

(i) Identify the metal **M**. Give the oxidation state of **M** in complex ion **P**.

(ii) Give the value of n in $M(NO_3)_n$.

(iii) Write the complete electronic configuration of **M** in complex ion **P**.

(iv) Write the chemical formulae of **P, Q, R, S, T** and **U**.

(v) Give the IUPAC names of **P, R, S, T** and **U**.

(vi) What is the colour of **P**?

(vii) What would you expect to observe in I and II given below?

I. When H_2S gas is passed into an acidic solution containing **P** at room temperature

II. When the mixture obtained in I above is heated with dilute HNO_3 after the removal of dissolved H_2S

(viii) Briefly describe a method with the aid of balanced chemical equations for determining the concentration of M^{n+} present in an aqueous solution, using the following chemicals.

KI, $Na_2S_2O_3$ and starch. **(75 *marks*)**

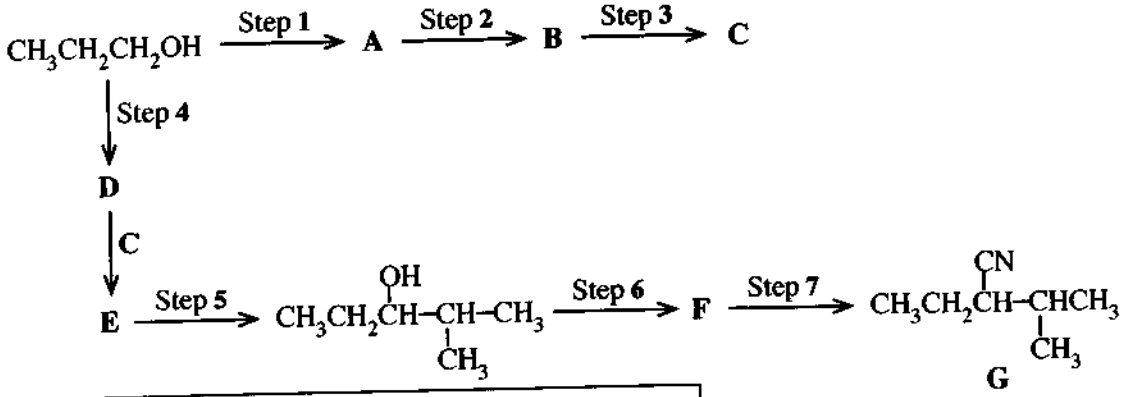
[*see page twelve*

PART C — ESSAY

Answer two questions only. (Each question carries **150** *marks.)*

8. (*a*) (i) Given below is a reaction scheme for the synthesis of compound **G** using $CH_3CH_2CH_2OH$ as the only organic starting compound.

Complete the reaction scheme by drawing the structures of compounds **A, B, C, D, E** and **F** and writing the appropriate reagents for steps **1 – 7**, selected only from those given in the list.



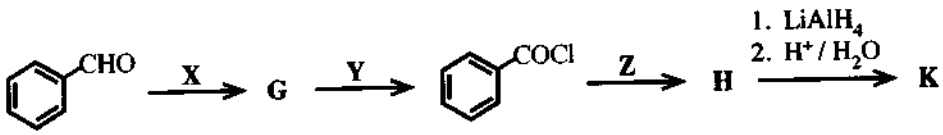
List of Reagents

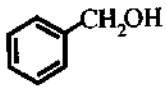
HBr, PBr_3, pyridiniumchlorochromate (PCC), Mg / dry ether, KCN, conc. H_2SO_4, dil. H_2SO_4

(52 *marks*)

(ii) Consider the following series of reactions.

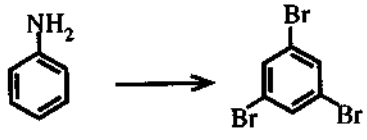
Draw the structures of compounds **G, H** and **K**. Give the reagents **X, Y** and **Z**.



Note that **K** gives benzyl alcohol () when reacted with $NaNO_2$/ dil. HCl.

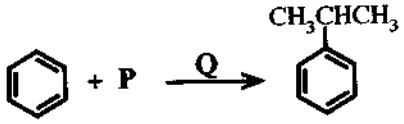
(24 *marks*)

(*b*) (i) Show how the following conversion could be carried out in **not more than three** steps.



(20 *marks*)

(ii) Consider the following reaction.



Identify the chemical substances **P** and **Q** necessary to carry out this reaction.
Write the mechanism of this reaction. **(20 *marks*)**

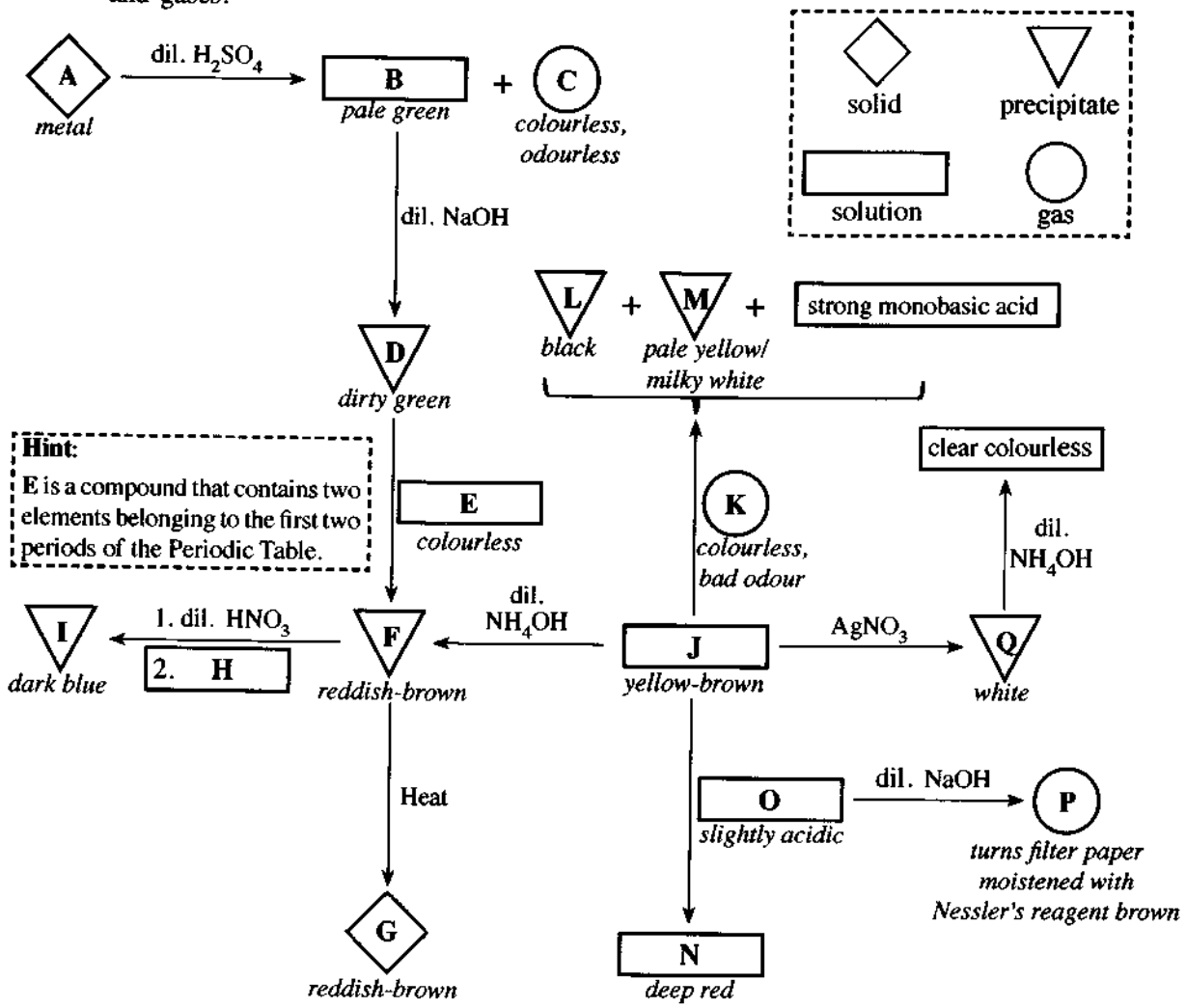
(*c*) (i) Explain why phenol is more reactive in electrophilic substitution reactions than benzene, by considering their resonance hybrids.

(ii) Illustrate the difference in reactivity between phenol and benzene as given in (i) above by means of a suitable reaction.

(iii) Draw the structure(s) of product(s) you described in the reaction in (ii) above. **(34 *marks*)**

[*see page thirteen*

9. (*a*) (i) Write the chemical formulae of the substances **A – Q** given in the flow chart below.
(**Note**: Chemical equations and reasons are not expected for the identification of substances **A – Q**.)
The symbols given in the box (dash lines) are used to represent solids, precipitates, solutions and gases.



(ii) Write the complete electronic configuration of **A**.

(iii) State the function of **E** in the conversion of **D** to **F**. Give the relevant balanced chemical equations for the stated function. ***(75 marks)***

(*b*) The solid **X** contains only Cu_2S and CuS. The following procedure was used to determine the percentage of Cu_2S in **X**.

Procedure

A 1.00 g portion of solid **X** was treated with 100.00 cm^3 of 0.16 mol dm^{-3} $KMnO_4$ in dilute H_2SO_4 medium. This reaction gave Mn^{2+}, Cu^{2+} and SO_4^{2-} as products. Thereafter, the excess $KMnO_4$ in this solution was titrated with 0.15 mol dm^{-3} Fe^{2+} solution. The volume required for the titration was 35.00 cm^3.

(i) Write the balanced ionic equations for the reactions taking place in the above procedure.

(ii) Based on the answers to (i) above, determine the molar ratio between,

I. Cu_2S and $KMnO_4$

II. CuS and $KMnO_4$

III. Fe^{2+} and $KMnO_4$

(iii) Calculate the percentage by weight of Cu_2S in **X**. (Cu = 63.5, S = 32) ***(75 marks)***

[*see page fourteen*

10. (*a*) The following questions are based on the properties of titanium dioxide (TiO_2) and its manufacture carried out by the "Chloride Process".

(i) Name the raw materials used in this process.

(ii) Briefly describe the manufacturing process of TiO_2 giving balanced chemical equations where applicable.

(iii) State **three** properties of TiO_2 and give one use each, relevant to each property.

(iv) If you were to consider establishing a TiO_2 manufacturing plant in Sri Lanka, state **three** requirements that need to be fulfilled.

(v) Does the manufacturing process described in (ii) above contribute to global warming? Justify your answer. **(50 *marks*)**

(*b*) Currently, global warming due to change in greenhouse effect is significantly greater than that before the industrial revolution.

(i) Explain briefly what is meant by greenhouse effect.

(ii) Identify the major environmental problem that occurs due to global warming.

(iii) State **two main** natural gases that contribute to global warming.

(iv) Explain briefly how microorganisms contribute to the release of the gases you stated in (iii).

(v) In addition to the gases you stated in (iii), name **two** classes of synthetic volatile compounds that directly contribute to the global warming, and selecting one compound from each class, draw their structures.

(vi) Select **one** class of compounds from the two classes you stated in (v) that contributes to the catalytic degradation of ozone in the upper atmosphere.

(vii) The slow down of industrial activities due to the Covid-19 pandemic temporarily eased the global environmental issues in many countries. Justify this statement by using **two** main global environmental issues you have learnt. **(50 *marks*)**

(*c*) The following questions are based on the polymers given below.

Polyvinyl chloride (PVC), Polyethylene (PE), Polystyrene (PS), Bakelite, Nylon 6.6, Polyethylene terephthalate (PET), Gutta percha

(i) Draw the repeating units of **four** of the above polymers.

(ii) Categorize each of the above seven (7) polymers as either,

I. natural or synthetic polymers.

II. addition or condensation polymers.

(iii) Name the **two** monomers used in the formation of bakelite.

(iv) Polymers can be grouped into two categories based on their thermal properties. State these **two** categories. Write to which of these categories PVC and bakelite belong.

(v) Give **one** use each for **three** of the polymers given in the above list. **(50 *marks*)**

* * *

[*see page fifteen*

නව නිර්දේශය/புதிய பாடத்திட்டம்/New Syllabus

NEW

අධ්‍යයන පොදු සහතික පත්‍ර (උසස් පෙළ) විභාගය, 2020
கல்விப் பொதுத் தராதரப் பத்திர (உயர் தர)ப் பரீட்சை, 2020
General Certificate of Education (Adv. Level) Examination, 2020

රසායන විද්‍යාව	II
இரசாயனவியல்	II
Chemistry	II

02 E II

පැය තුනයි மூன்று மணித்தியாலம் ***Three hours***	අමතර කියවීම් කාලය - මිනිත්තු 10 යි மேலதிக வாசிப்பு நேரம் - 10 நிமிடங்கள் **Additional Reading Time - 10 minutes**

Use **additional reading time** to go through the question paper, select the questions and decide on the questions that you give priority in answering.

Index No. : .....

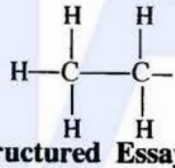
* *A Periodic Table is provided on page 15.*

* ***Use of calculators is not allowed.***

* *Universal gas constant,* $R = 8.314\ J\ K^{-1}\ mol^{-1}$

* *Avogadro constant,* $N_A = 6.022 \times 10^{23}\ mol^{-1}$

* **In answering this paper, you may represent alkyl groups in a condensed manner.**

Example:  group may be shown as CH_3CH_2-

❑ **PART A — Structured Essay (pages 02 - 08)**

* *Answer **all** the questions on the question paper itself.*

* *Write your answer in the space provided for each question. Please note that the space provided is sufficient for the answer and that extensive answers are not expected.*

❑ **PART B and PART C — Essay (pages 09 - 14)**

* *Answer **four** questions selecting **two** questions from each part. Use the papers supplied for this purpose.*

* *At the end of the time allotted for this paper, tie the answers to the three Parts **A**, **B** and **C** together so that Part **A** is on top and hand them over to the Supervisor.*

* *You are permitted to remove **only** Parts **B** and **C** of the question paper from the Examination Hall.*

For Examiner's Use Only

Part	Question No.	Marks
A	1	
	2	
	3	
	4	
B	5	
	6	
	7	
C	8	
	9	
	10	
Total		

Total

In Numbers	
In Letters	

Code Numbers

Marking Examiner 1	
Marking Examiner 2	
Checked by :	
Supervised by :	

[see page two

PART A — STRUCTURED ESSAY

*Answer **all four** questions on this paper itself. (Each question carries **100** marks.)*

Do not write in this column.

1. (*a*) Write the answers to the questions given below on the dotted lines.

(i) Of the three ions Na^+, Mg^{2+} and F^-, which one has the **smallest** ionic radius? .....

(ii) Of the three elements C, N and O, which one has the **highest** second ionization energy? .....

(iii) Of the three compounds H_2O, HOCl and OF_2, which one has the **most** electronegative oxygen atom? .....

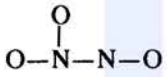
(iv) Of the three elements Be, C and N, which one will liberate energy when an electron is added to its atom [$Y(g) + e \rightarrow Y^-(g)$; Y = Be, C, N] in the gaseous state? .....

(v) Of the three ionic compounds NaF, KF and KBr, which one has the **highest** solubility in water? .....

(vi) Of the three compounds HCHO, CH_3F and H_2O_2, which one has the **strongest** intermolecular forces? .....

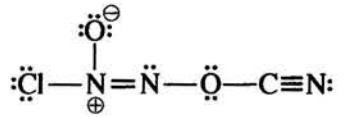
(24 *marks*)

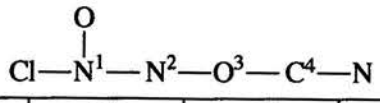
(*b*) (i) Draw the **most** acceptable Lewis dot-dash structure for the ion, $N_2O_3^{2-}$. Its skeleton is given below.



(ii) Draw **three** more Lewis dot-dash structures (resonance structures) for this ion. Indicate the relative stabilities of the structures drawn by you, when compared with the most acceptable structure drawn in (i) above, by writing '**less stable**' or '**unstable**' under these structures.

(iii) Complete the given table based on the Lewis dot-dash structure and its labelled skeleton given below.





	N^1	N^2	O^3	C^4
VSEPR pairs around the atom				
electron pair geometry around the atom				
shape around the atom				
hybridization of the atom				

[*see page three*

Do not write in this column.

- Parts (iv) to (vii) are based on the Lewis dot-dash structure given in part (iii) above. Labelling of atoms is as in part (iii).

(iv) Identify the atomic / hybrid orbitals involved in the formation of σ bonds between the two atoms given below.

I.	$Cl—N^1$	Cl	N^1	.....
II.	$N^1—O$	N^1	O	.....
III.	$N^1—N^2$	N^1	N^2	.....
IV.	$N^2—O^3$	N^2	O^3	.....
V.	$O^3—C^4$	O^3	C^4	.....
VI.	$C^4—N$	C^4	N	.....

(v) Identify the atomic orbitals involved in the formation of π bonds between the two atoms given below.

I.	$N^1—N^2$	N^1	N^2	.....
II.	$C^4—N$	C^4	N	.....
		C^4	N	.....

(vi) State the approximate bond angles around N^1, N^2, O^3 and C^4 atoms.

N^1....., N^2 ....., O^3 ....., C^4.....

(vii) Arrange the atoms N^1, N^2, O^3 and C^4 in the **increasing** order of electronegativity.

..... < ..... < ..... < ..... **(56 *marks*)**

(*c*) Consider the following information.

I. The atoms **A** and **B** combine to form a heterodiatomic molecule **AB** that has a σ bond. This is represented as **A–B**.

II. The electronegativity of **A** is less than that of **B** ($X_A < X_B$).
X = electronegativity of the atom

III. The inter-nuclear distance between **A** and **B** atoms (d_{A-B}) of the **AB** molecule is given by the following equation.

$d_{A-B} = r_A + r_B - c(X_B - X_A)$

r = atomic radius, c = 9 pm

Note: d and r are measured in picometres (pm). (1 pm = 10^{-12} m)

Based on the above information, answer the following questions.

(i) What is the name used to identify the type of σ bond between **A** and **B**?

.....

(ii) Show how fractional charges ($\delta+$ and $\delta-$) are located in the molecule **AB**.

.....

(iii) Write the equation to calculate the dipole moment (μ) of molecule **AB** and show its direction.

[*see page four*

Do not write in this column.

(iv) Calculate the percentage of ionic character of the H–F bond in the HF molecule using the data given below.

Inter-nuclear distance of H_2 (d_{H-H})	= 74 pm	Electronegativity of F	= 4.0
Inter-nuclear distance of F_2 (d_{F-F})	= 144 pm	Dipole moment of HF	= 6.0×10^{-30} C m
Electronegativity of H	= 2.1	Charge of an electron	= 1.6×10^{-19} C

100

(20 *marks*)

2. (*a*) **A, B, C** and **D** are chlorides of *p*-block elements. These elements have atomic numbers less than 20. A description of the products ($P_1 - P_9$) formed when **A** is reacted with a limited amount of water and **B, C** and **D** are reacted with excess water are given below.

Compound	Description of products	
A	P_1	a compound with a covalent network structure
	P_2	a strong monobasic acid
B	P_3	a gas that turns red litmus blue
	P_4	a compound with bleaching properties
C	P_5	a tribasic acid
	P_6	a strong monobasic acid
D	P_7	a gas that turns acidic $KMnO_4$ solution colourless
	P_8	a colloidal solid
	P_9	a strong monobasic acid

(i) Identify **A, B, C** and **D** (give the chemical formulae).

A: ..... **B:** ..... **C:** ..... **D:** .....

(ii) Give balanced chemical equations for the reactions of **A, B, C** and **D** with water to give products P_1 to P_9.

.....

.....

.....

.....

[*see page five*

Do not write in this column.

(iii) Write balanced chemical equations for the following reactions.

I. $\mathbf{P_1}$ with NaOH(aq)

.....

II. $\mathbf{P_3}$ with Mg

.....

III. $\mathbf{P_7}$ with acidic $K_2Cr_2O_7$

.....

(50 *marks*)

(*b*) A student is provided with bottles labelled **P, Q, R, S, T** and **U** containing aqueous solutions of $Al_2(SO_4)_3$, H_2SO_4, $Na_2S_2O_3$, $BaCl_2$, $Pb(Ac)_2$ and KOH **(not in order)**. Some useful observations for their identification on mixing two solutions at a time are given below. (Ac - Acetate ion)

	Solutions mixed	Observations
I	**T + R**	a clear colourless solution
II	**P + R**	a white precipitate
III	**T + S**	a gelatinous white precipitate
IV	**U + R**	a white precipitate
V	**P + Q**	a white precipitate, turns black on heating
VI	**P + U**	a white precipitate, dissolves on heating

(i) Identify **P** to **U**.

P: ..... **Q**: ..... **R**: .....

S: ..... **T**: ..... **U**: .....

(ii) Give balanced chemical equations for each of the reactions **I** to **VI**.

I: .....

II: .....

III: .....

IV: .....

V: formation of white precipitate: .....

turning black on heating: .....

VI: .....

(**Note**: indicate precipitates as ↓)

(50 *marks*)

100

3. (*a*) A saturated aqueous solution of a sparingly soluble salt $AB_2(s)$ was prepared by stirring an excess amount of $AB_2(s)$ in 1.0 dm^3 of distilled water at 25 °C. The amount of $A^{2+}(aq)$ ions present in this saturated aqueous solution was found to be 2.0×10^{-3} mol.

(i) Write the equilibrium related to the dissolution of $AB_2(s)$ in the above system at 25 °C.

.....

(ii) Write the expression for the equilibrium constant for the equilibrium written in (i) above at 25 °C.

.....

.....

[*see page six*

Do not write in this column.

(iii) Calculate the value of the equilibrium constant stated in (ii) above at 25 °C.

.....

.....

.....

.....

.....

.....

(iv) Another saturated aqueous solution of AB_2 was prepared by stirring an excess amount of $AB_2(s)$ in 2.0 dm^3 of distilled water at 25 °C. Giving reasons, predict the value of the equilibrium constant for this system.

.....

.....

(v) A small amount of the strong electrolyte NaB(s) is added to a saturated aqueous solution of AB_2 at 25 °C. Giving reasons, predict whether the concentration of $A^{2+}(aq)$ is increased or decreased.

.....

.....

.....

(60 *marks*)

(*b*) In an aqueous solution, propanoic acid (C_2H_5COOH) ionizes as given below.

$C_2H_5COOH(aq) + H_2O(l) \rightleftharpoons C_2H_5COO^-(aq) + H_3O^+(aq)$

At 25 °C, K_a (propanoic acid) = 1.0×10^{-5}

(i) Write the expression for the equilibrium constant for the above reaction at 25 °C.

.....

.....

(ii) 100.0 cm^3 of an aqueous solution of $C_2H_5COOH(aq)$ was prepared by dissolving 0.74 cm^3 of C_2H_5COOH in distilled water at 25 °C. Calculate the pH of the solution at 25 °C.
(C = 12; O = 16; H = 1; consider the density of C_2H_5COOH as 1.0 g cm^{-3})

(40 *marks*)

100

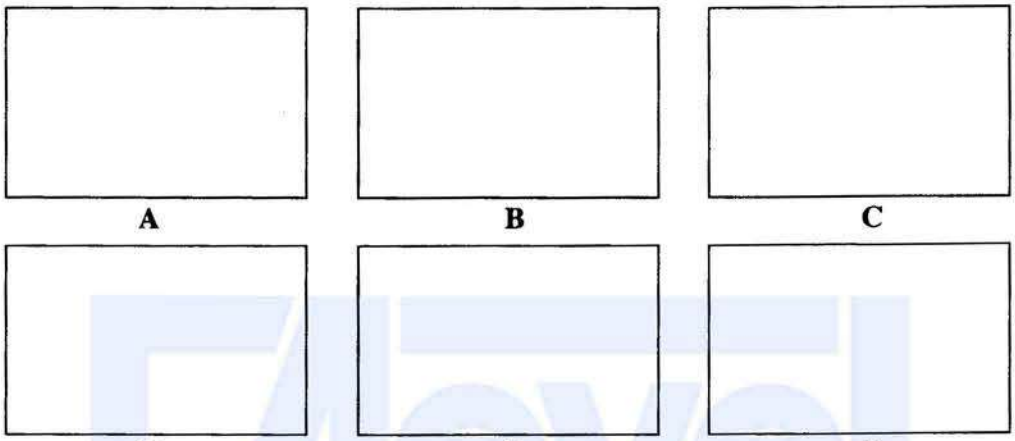
[*see page seven*

Do not write in this column.

4. (*a*) **A, B, C** and **D** are structural isomers having the molecular formula C_6H_{10}. None of them show optical isomerism. All four isomers, **A, B, C** and **D** when treated with $HgSO_4$/dil. H_2SO_4 give products which react with 2,4-dinitrophenylhydrazine (2,4-DNP) to give coloured precipitates.

Only **A** gives a precipitate with ammonical $AgNO_3$. **A** has only one position isomer, which is **B**. **B** is a chain isomer of **C**. **C** reacts with $HgSO_4$/dil. H_2SO_4 to give two products **E** and **F**. **D** reacts with $HgSO_4$/dil. H_2SO_4 to give only one product, which is **E**.

(i) Draw the structures of **A, B, C, D, E** and **F** in the boxes given below.



A **B** **C**

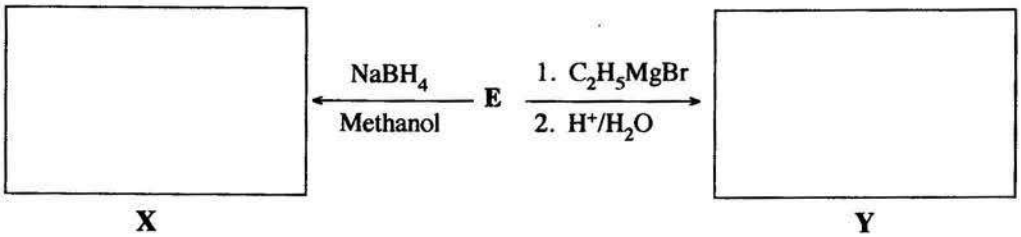
D **E** **F**

(ii) Which of the compounds **A, B, C** and **D** gives a product that does not show diastereoisomerism when reacted separately with H_2 / Pd-$BaSO_4$ / quinoline?

(iii) Draw, in the box given below, the structure of the product **G** obtained when **A** is reacted with excess HBr.

G

(iv) Draw the structures of products **X** and **Y** obtained in the following reactions of **E**, in the appropriate boxes.



X **Y**

Name a test to distinguish between **X** and **Y**.

.....

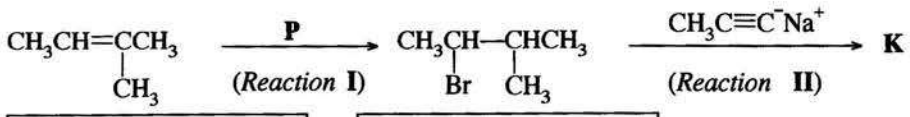
(60 *marks*)

[*see page eight*

Do not write in this column.

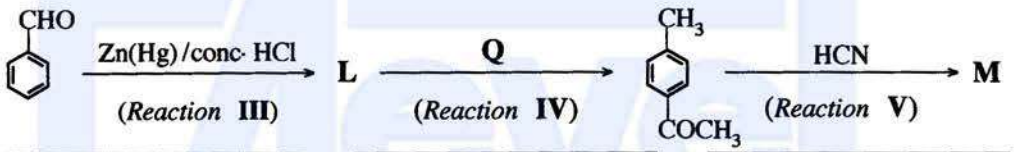
(*b*) (i) Complete the following three reaction sequences by drawing structures of compounds **K, L** and **M** and giving the reagents/catalysts **P, Q** and **R** in the boxes given below.

Sequence 1:



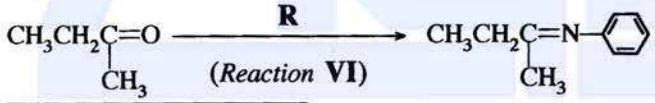
P **K**

Sequence 2:



L **Q** **M**

Sequence 3:



R

(30 *marks*)

(ii) Selecting from the reactions **I – VI**, give **one (01)** example for each of the following types of reactions.

Nucleophilic addition .....

Nucleophilic substitution ..... **(10 *marks*)**

100

* *

[*see page nine*

නව නිර්දේශය/புதிய பாடத்திட்டம்/New Syllabus

NEW

අධ්‍යයන පොදු සහතික පත්‍ර (උසස් පෙළ) විභාගය, 2020
கல்விப் பொதுத் தராதரப் பத்திர (உயர் தர)ப் பரீட்சை, 2020
General Certificate of Education (Adv. Level) Examination, 2020

රසායන විද්‍යාව	II
இரசாயனவியல்	II
Chemistry	**II**

02 E II

* Universal gas constant $R = 8.314\ J\ K^{-1}\ mol^{-1}$

* Avogadro constant $N_A = 6.022 \times 10^{23}\ mol^{-1}$

PART B — ESSAY

*Answer **two** questions only. (Each question carries **150 marks**.)*

5. (*a*) A compound $XY_2Z_2(g)$ undergoes dissociation when heated to temperatures above 300 K as given below.

$$XY_2Z_2(g) \overset{\Delta}{\rightleftharpoons} XY_2(g) + Z_2(g)$$

A sample of 7.5 g of $XY_2Z_2(g)$ was placed in an evacuated 1.00 dm^3 rigid-closed container and the temperature was raised to 480 K.
Molar mass of $XY_2Z_2(g)$ is 150 g mol^{-1}. Use the approximate value of 4000 J mol^{-1} for RT at 480 K. Assume ideal gas behaviour for all gases.

(i) Calculate the number of moles of $XY_2Z_2(g)$ in the container before dissociation.

(ii) When the above system reaches equilibrium at 480 K, the total number of moles in the container was found to be 7.5×10^{-2} mol. Calculate the number of moles of $XY_2Z_2(g)$, $XY_2(g)$ and $Z_2(g)$ in the equilibrium mixture at 480 K.

(iii) Calculate the equilibrium constant K_c for the above reaction at 480 K.

(iv) Calculate K_p for the equilibrium at 480 K. **(75 *marks*)**

(*b*) For the reaction $XY_2Z_2(g) \rightarrow XY_2(g) + Z_2(g)$ described in (a), Gibbs free energies (G) at 480 K for $XY_2Z_2(g)$, $XY_2(g)$ and $Z_2(g)$ are -60 kJ mol^{-1}, -76 kJ mol^{-1} and -30 kJ mol^{-1}, respectively.

(i) Calculate ΔG (in kJ mol^{-1}) for the reaction at 480 K.

(ii) The magnitude of ΔS of the above reaction is 150 J K^{-1} mol^{-1} at 480 K. Calculate ΔH for the reaction at 480 K by using the appropriate sign (– or +) of ΔS.

(iii) By using the sign (– or +) of ΔH obtained in (ii), explain whether this reaction is exothermic or endothermic.

(iv) Deduce the enthalpy difference for the formation of $XY_2Z_2(g)$ from $XY_2(g)$ and $Z_2(g)$ at 480 K.

(v) If the bond enthalpy of the X–Z bond in $XY_2Z_2(g)$ is +250 kJ mol^{-1}, calculate the bond enthalpy of the Z–Z bond.
(Assume that $XY_2Z_2(g)$ has the structure $Z-\overset{\overset{Y}{\|}}{\underset{\underset{Y}{\|}}{X}}-Z$)

(vi) If liquid XY_2Z_2 is used instead of gaseous XY_2Z_2, giving reasons, explain whether the value of ΔH obtained for the reaction $XY_2Z_2(l) \rightarrow XY_2(g) + Z_2(g)$ is equal to, or higher or lower than ΔH obtained in (ii). **(75 *marks*)**

[*see page ten*

6. (*a*) Consider the reaction given below occurring in a closed container at a given temperature *T*.

$$2N_2O_5(g) \rightarrow 4NO_2(g) + O_2(g)$$

(i) Write **three** expressions for the rate of reaction relevant to each of the compounds appearing in the reaction.

(ii) This reaction was carried out at temperature *T* with an initial concentration of 0.10 mol dm^{-3} of $N_2O_5(g)$. It was found that 40% of the initial amount was decomposed after a period of 400 s.

I. Calculate the average rate of decomposition of $N_2O_5(g)$ in this time interval.

II. Calculate average rates of formation of $NO_2(g)$ and $O_2(g)$.

(iii) In another experiment, initial rates were measured for this reaction at 300 K and the results are given below.

$[N_2O_5(g)]$ / mol dm^{-3}	0.01	0.02	0.03
Initial rate / mol dm^{-3} s^{-1}	6.930×10^{-5}	1.386×10^{-4}	2.079×10^{-4}

Derive the rate law for the reaction at 300 K.

(iv) Another experiment was carried out at 300 K with an initial concentration of 0.64 mol dm^{-3} of $N_2O_5(g)$. It was found that the concentration of $N_2O_5(g)$ which remained after a period of 500 s was 2.0×10^{-2} mol dm^{-3}.

I. Calculate the half-life ($t_{1/2}$) of the reaction at 300 K.

II. Calculate the rate constant of the reaction at 300 K.

(v) This reaction proceeds through a mechanism involving the following elementary steps.

Step 1 : $N_2O_5(g) \rightleftharpoons NO_3(g) + NO_2(g)$: Fast

Step 2 : $NO_3(g) + NO_2(g) \rightarrow 2NO_2(g) + O(g)$: Slow

Step 3 : $N_2O_5(g) + O(g) \rightarrow 2NO_2(g) + O_2(g)$: Fast

Show that the above mechanism is consistent with the rate law of the reaction. **(80 *marks*)**

(*b*) An ideal binary-liquid mixture was prepared by mixing two liquids of **A** and **B** in a closed evacuated container at temperature *T*. After establishing the equilibrium at temperature *T*, partial pressures of **A** and **B** in the vapour phase are P_A and P_B, respectively. At temperature *T*, the saturated vapour pressures of **A** and **B** are P_A° and P_B°, respectively. Mole fractions of **A** and **B** in solution are X_A and X_B, respectively.

(i) Show that $P_A = P_A^\circ X_A$

(Consider that the rates of vaporization and condensation are equal at equilibrium.)

(ii) In the above system at 300 K, the total pressure was 5.0×10^4 Pa. The saturated vapour pressures of pure **A** and **B** at 300 K, are 7.0×10^4 Pa and 3.0×10^4 Pa, respectively.

I. Calculate the mole fraction of **A** in the liquid phase of the equilibrium mixture.

II. Calculate the vapour pressure of **A** in the equilibrium mixture.

(70 *marks*)

[*see page eleven*

7. (*a*) (i) To compare the properties of Electrolytic and Galvanic cells, copy and complete the following table using the given terms.

Terms: anode, cathode, positive, negative, spontaneous, non-spontaneous.

		Electrolytic cell	Galvanic cell
A.	Oxidation half-reaction takes place at		
B.	Reduction half-reaction takes place at		
C.	Sign of E°_{cell}		
D.	Electron flow	From ..... to	From ..... to .....
E.	Spontaneity of the cell reaction		

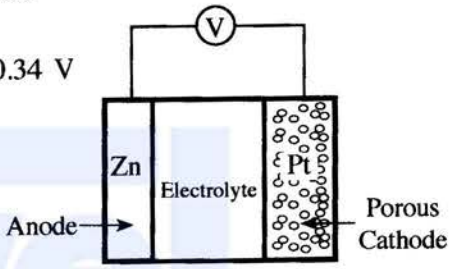
(ii) An electrochemical cell was constructed at 300 K by using a Zn(s) anode, an aqueous alkaline electrolyte and a porous Pt cathode which facilitates the collection of oxygen $O_2(g)$ from air as shown below. As the cell operates ZnO(s) is produced.

You are given that

$E^\circ_{ZnO(s)\,|\,Zn(s)\,|\,OH^-(aq)} = -1.31$ V and $E^\circ_{O_2(g)\,|\,OH^-(aq)} = +0.34$ V

Zn = 65 g mol^{-1}, O = 16 g mol^{-1} and

1 F = 96,500 C



I. Write the half-reactions occurring at anode and cathode.

II. Write the overall cell reaction.

III. Calculate the cell potential E°_{cell} at 300 K.

IV. State the direction of migration of $OH^-(aq)$ ions between the electrodes.

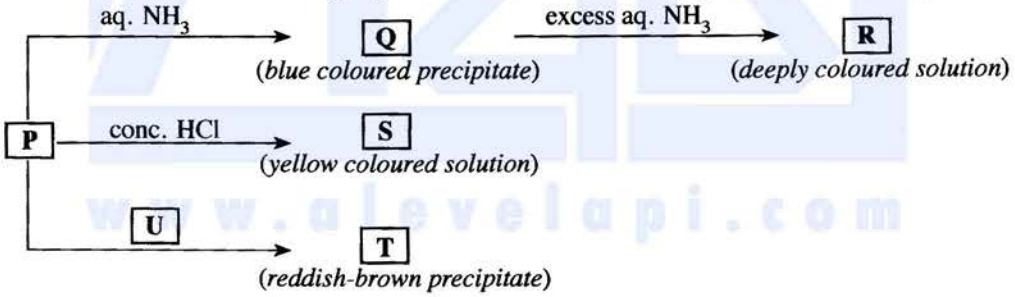
V. When the cell operates for a period of 800 s at 300 K, 2 mol of $O_2(g)$ are consumed.

A. Calculate the number of moles of electrons passing through the cell.

B. Calculate the mass of ZnO(s) formed.

C. Calculate the current passing through the cell. **(75 *marks*)**

(*b*) A coloured complex ion **P** is formed when the salt $\mathbf{M(NO_3)_n}$ is dissolved in distilled water. **M** is a transition element belonging to the 3*d* block. **P** undergoes the following reactions.



T and **U** are coordination compounds each containing four elements. **P, R** and **S** are complex ions.

(i) Identify the metal **M**. Give the oxidation state of **M** in complex ion **P**.

(ii) Give the value of n in $\mathbf{M(NO_3)_n}$.

(iii) Write the complete electronic configuration of **M** in complex ion **P**.

(iv) Write the chemical formulae of **P, Q, R, S, T** and **U**.

(v) Give the IUPAC names of **P, R, S, T** and **U**.

(vi) What is the colour of **P**?

(vii) What would you expect to observe in I and II given below?

I. When H_2S gas is passed into an acidic solution containing **P** at room temperature

II. When the mixture obtained in I above is heated with dilute HNO_3 after the removal of dissolved H_2S

(viii) Briefly describe a method with the aid of balanced chemical equations for determining the concentration of $\mathbf{M^{n+}}$ present in an aqueous solution, using the following chemicals.

KI, $Na_2S_2O_3$ and starch. **(75 *marks*)**

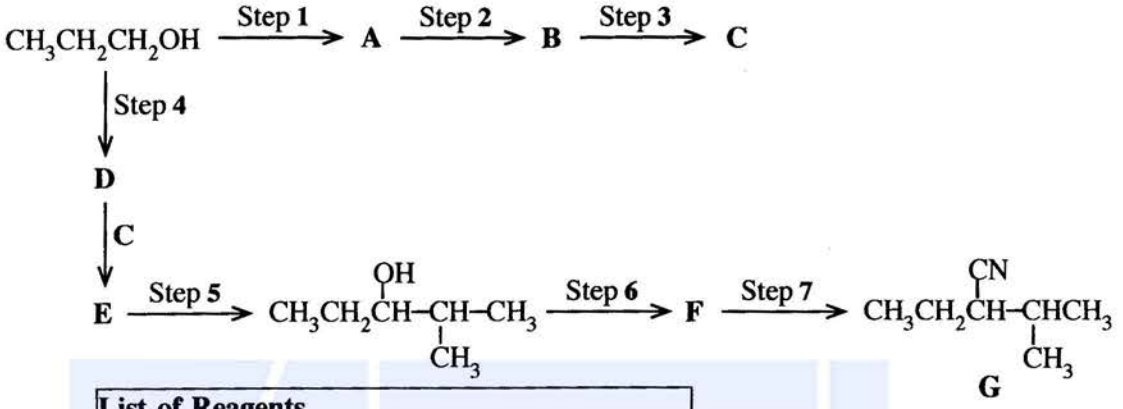
[*see page twelve*

PART C — ESSAY

*Answer **two** questions only. (Each question carries **150 marks**.)*

8. (*a*) (i) Given below is a reaction scheme for the synthesis of compound **G** using $CH_3CH_2CH_2OH$ as the only organic starting compound.

Complete the reaction scheme by drawing the structures of compounds **A, B, C, D, E** and **F** and writing the appropriate reagents for steps **1 – 7**, selected only from those given in the list.



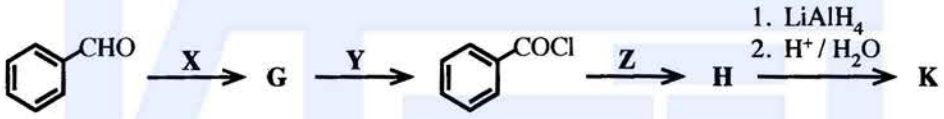
List of Reagents

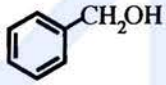
HBr, PBr_3, pyridiniumchlorochromate (PCC),
Mg / dry ether, KCN, conc. H_2SO_4, dil. H_2SO_4

(52 *marks*)

(ii) Consider the following series of reactions.

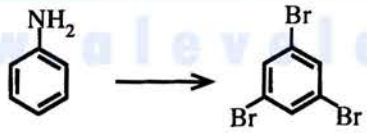
Draw the structures of compounds **G, H** and **K**. Give the reagents **X, Y** and **Z**.



Note that **K** gives benzyl alcohol () when reacted with $NaNO_2$/ dil. HCl.

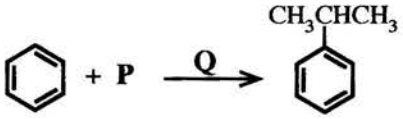
(24 *marks*)

(*b*) (i) Show how the following conversion could be carried out in **not more than three** steps.



(20 *marks*)

(ii) Consider the following reaction.



Identify the chemical substances **P** and **Q** necessary to carry out this reaction.
Write the mechanism of this reaction. **(20 *marks*)**

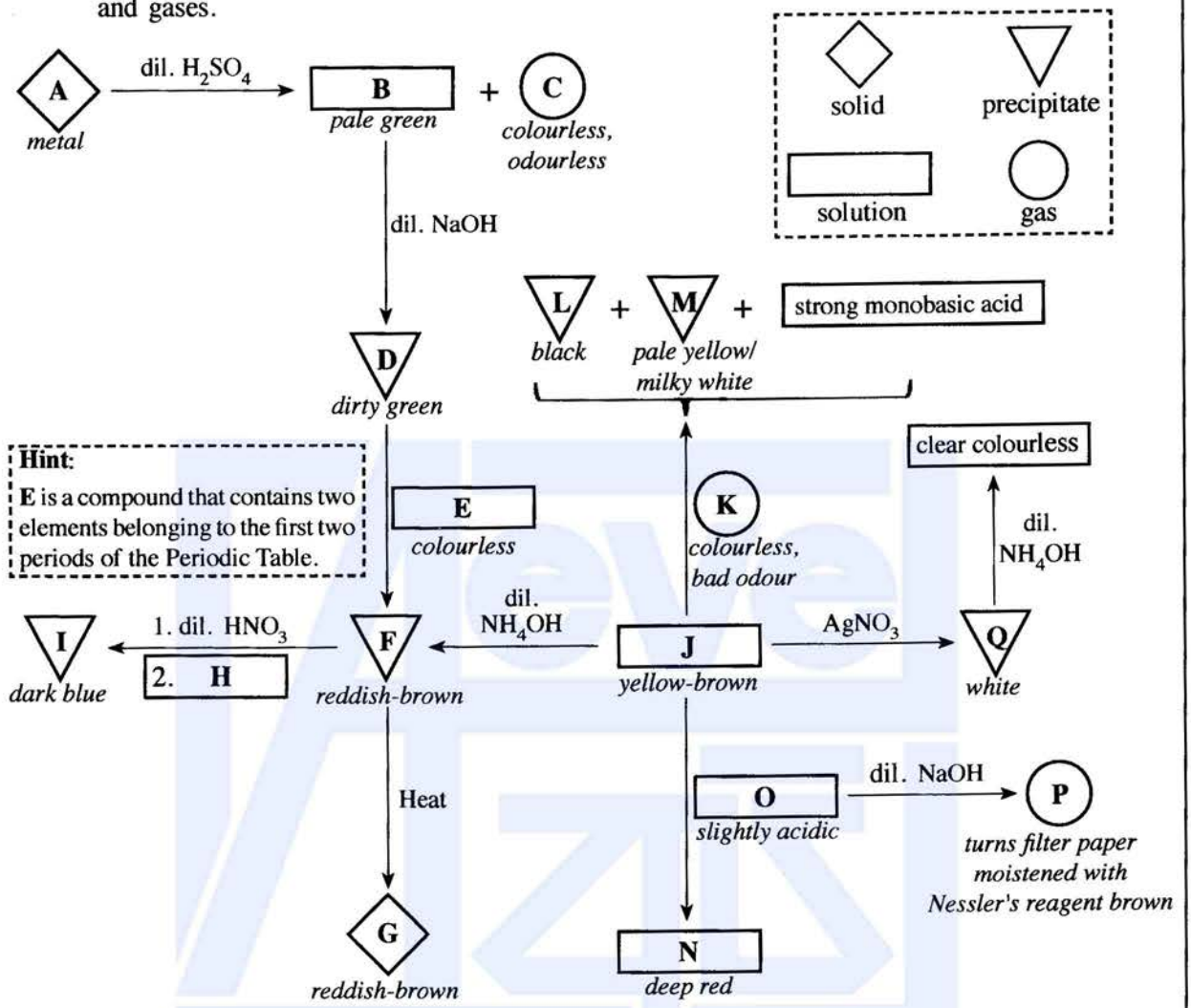
(*c*) (i) Explain why phenol is more reactive in electrophilic substitution reactions than benzene, by considering their resonance hybrids.

(ii) Illustrate the difference in reactivity between phenol and benzene as given in (i) above by means of a suitable reaction.

(iii) Draw the structure(s) of product(s) you described in the reaction in (ii) above. **(34 *marks*)**

[*see page thirteen*

9. (*a*) (i) Write the chemical formulae of the substances **A – Q** given in the flow chart below.
(**Note**: Chemical equations and reasons are not expected for the identification of substances **A – Q**.)
The symbols given in the box (dash lines) are used to represent solids, precipitates, solutions and gases.



(ii) Write the complete electronic configuration of **A**.

(iii) State the function of **E** in the conversion of **D** to **F**. Give the relevant balanced chemical equations for the stated function. **(75 *marks*)**

(*b*) The solid **X** contains only Cu_2S and CuS. The following procedure was used to determine the percentage of Cu_2S in **X**.

Procedure

A 1.00 g portion of solid **X** was treated with 100.00 cm^3 of 0.16 mol dm^{-3} $KMnO_4$ in dilute H_2SO_4 medium. This reaction gave Mn^{2+}, Cu^{2+} and SO_4^{2-} as products. Thereafter, the excess $KMnO_4$ in this solution was titrated with 0.15 mol dm^{-3} Fe^{2+} solution. The volume required for the titration was 35.00 cm^3.

(i) Write the balanced ionic equations for the reactions taking place in the above procedure.

(ii) Based on the answers to (i) above, determine the molar ratio between,

I. Cu_2S and $KMnO_4$
II. CuS and $KMnO_4$
III. Fe^{2+} and $KMnO_4$

(iii) Calculate the percentage by weight of Cu_2S in **X**. (Cu = 63.5, S = 32) **(75 *marks*)**

[*see page fourteen*

10. (*a*) The following questions are based on the properties of titanium dioxide (TiO_2) and its manufacture carried out by the "Chloride Process".

(i) Name the raw materials used in this process.

(ii) Briefly describe the manufacturing process of TiO_2 giving balanced chemical equations where applicable.

(iii) State **three** properties of TiO_2 and give one use each, relevant to each property.

(iv) If you were to consider establishing a TiO_2 manufacturing plant in Sri Lanka, state **three** requirements that need to be fulfilled.

(v) Does the manufacturing process described in (ii) above contribute to global warming? Justify your answer. **(50 *marks*)**

(*b*) Currently, global warming due to change in greenhouse effect is significantly greater than that before the industrial revolution.

(i) Explain briefly what is meant by greenhouse effect.

(ii) Identify the major environmental problem that occurs due to global warming.

(iii) State **two main** natural gases that contribute to global warming.

(iv) Explain briefly how microorganisms contribute to the release of the gases you stated in (iii).

(v) In addition to the gases you stated in (iii), name **two** classes of synthetic volatile compounds that directly contribute to the global warming, and selecting one compound from each class, draw their structures.

(vi) Select **one** class of compounds from the two classes you stated in (v) that contributes to the catalytic degradation of ozone in the upper atmosphere.

(vii) The slow down of industrial activities due to the Covid-19 pandemic temporarily eased the global environmental issues in many countries. Justify this statement by using **two** main global environmental issues you have learnt. **(50 *marks*)**

(*c*) The following questions are based on the polymers given below.

Polyvinyl chloride (PVC), Polyethylene (PE), Polystyrene (PS), Bakelite, Nylon 6.6, Polyethylene terephthalate (PET), Gutta percha

(i) Draw the repeating units of **four** of the above polymers.

(ii) Categorize each of the above seven (7) polymers as either,

I. natural or synthetic polymers.

II. addition or condensation polymers.

(iii) Name the **two** monomers used in the formation of bakelite.

(iv) Polymers can be grouped into two categories based on their thermal properties. State these **two** categories. Write to which of these categories PVC and bakelite belong.

(v) Give **one** use each for **three** of the polymers given in the above list. **(50 *marks*)**

* * *

[*see page fifteen*

ශ්‍රී ලංකා විභාග දෙපාර්තමේන්තුව

Department of Examinations – Sri Lanka

අ.පො.ස.(උ.පෙළ)විභාගය/G.C.E. (A/L)- 2020

නව නිර්දේශය/ New Syllabus

විෂයය අංකය Subject No: 02

විෂයය Subject: Chemistry

ලකුණු දීමේ පටිපාටිය/Marking Scheme

I පත්‍රය/Paper I

ප්‍රශ්න අංකය Question No.	පිළිතුරු අංකය Answer No.	ප්‍රශ්න අංකය Question No.	පිළිතුරු අංකය Answer No.	ප්‍රශ්න අංකය Question No.	පිළිතුරු අංකය Answer No.	ප්‍රශ්න අංකය Question No.	පිළිතුරු අංකය Answer No.	ප්‍රශ්න අංකය Question No.	පිළිතුරු අංකය Answer No.
01.	5	11.	2	21.	3	31.	5	41.	4
02.	3	12.	3	22.	4 or 5	32.	2	42.	1 or 2
03.	4	13.	3	23.	1	33.	5	43.	3
04.	2	14.	2	24.	ALL	34.	4 or 5	44.	4
05.	ALL	15.	ALL	25.	ALL	35.	1	45.	5
06.	1	16.	3	26.	1	36.	5	46.	1
07.	2	17.	1	27.	5	37.	3 or 5	47.	4
08.	3	18.	1	28.	5	38.	4	48.	1
09.	4	19.	2	29.	4	39.	4	49.	3
10.	2	20.	2	30.	2	40.	5	50.	3

✪විශේෂ උපදෙස්/ Special Instructions:

එක් පිළිතුරකට ලකුණු 01 බැගින්/ 01 Mark for each question

මුළු ලකුණු/ Total Marks 01 × 50 = 50

PART A — STRUCTURED ESSAY

Answer all four questions on this paper itself. (Each question carries 10 marks.)

1. (*a*) Write the answers to the questions given below on the dotted lines.

(i) Of the three ions Na^+, Mg^{2+} and F^-, which one has the **smallest** ionic radius? ..... Mg^{2+} .....

(ii) Of the three elements C, N and O, which one has the **highest** second ionization energy? ..... O .....

(iii) Of the three compounds H_2O, HOCl and OF_2, which one has the **most** electronegative oxygen atom? ..... OF_2 .....

(iv) Of the three elements Be, C and N, which one will liberate energy when an electron is added to its atom [$Y(g) + e \rightarrow Y^-(g)$; Y = Be, C, N] in the gaseous state? ..... C .....

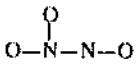
(v) Of the three ionic compounds NaF, KF and KBr, which one has the **highest** solubility in water? ..... KF or KBr .....

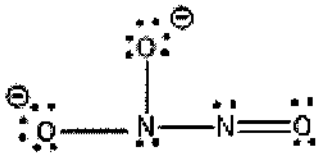
(vi) Of the three compounds HCHO, CH_3F and H_2O_2, which one has the **strongest** intermolecular forces? ..... H_2O_2 .....

(04 marks X 6 = 24)

1(a): 24 marks

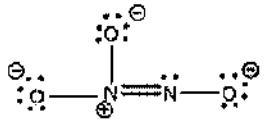
(*b*) (i) Draw the **most** acceptable Lewis structure for the ion, $N_2O_3^{2-}$. Its skeleton is given below.





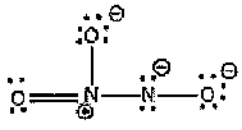
(06)

(ii) Draw **three** more Lewis structures (resonance structures) for this ion. Indicate the relative stabilities of the structures drawn by you, when compared with the most acceptable structure drawn in (i) above, by writing **'less stable'** or **'unstable'** under these structures.



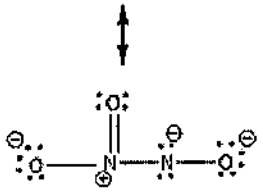
(02)

(less stable) **(01)**



(02)

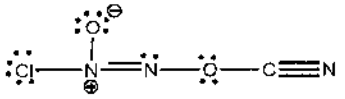
(unstable) **(01)**

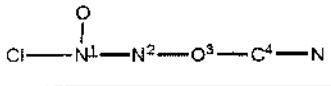


(02)

(unstable) **(01)**

(iii) Complete the given table based on the Lewis structure and its labelled skeleton given below.





	N^1	N^2	O^3	C^4
VSEPR pairs around the atom	3	3	4	2
electron pair geometry around the atom	trigonal planar	trigonal planar	tetrahedral	linear
shape around the atom	trigonal planar	angular/ V	angular/ V	linear
hybridization of the atom	sp^2	sp^2	sp^3	sp

(01 X 16 = 16)

- Parts (iv) to (vii) are based on the Lewis structure given in part (iii) above. Labelling of atoms is as in part (iii).

(iv) Identify the atomic/hybrid orbitals involved in the formation of σ bonds between the two atoms given below.

I. Cl—N^1	Cl	$3p$ **OR** sp^3	N^1	sp^2
II. N^1—O	N^1	sp^2	O	$2p$ **OR** sp^3
III. N^1—N^2	N^1	sp^2	N^2	sp^2
IV. N^2—O^3	N^2	sp^2	O^3	sp^3
V. O^3—C^4	O^3	sp^3	C^4	sp
VI. C^4—N	C^4	sp	N	$2p$ **OR** sp

(01 X 12 = 12)

(v) Identify the atomic orbitals involved in the formation of π bonds between the two atoms given below.

I. N^1—N^2	N^1	N^1, $2p$	N^2	N^2, $2p$
II. C^4—N	C^4	C^4, $2p$	N	N, $2p$
	C^4	C^4, $2p$	N	N, $2p$

(01 X 6 = 06)

(vi) State the approximate bond angles around N^1, N^2, O^3 and C^4 atoms.

N^1 $120° \pm 1$, N^2 $115°$-$118°$ O^3 $104° \pm 1$, C^4 $180° \pm 1$

(01 X 4 = 04)

(vii) Arrange the atoms N^1, N^2, O^3 and C^4 in the **increasing** order of electronegativity.

C^4 < N^2 < N^1 < O^3 **(03)**

1(b): 56 marks

(c) Consider the following information.

I. The atoms **A** and **B** combine to form a heterodiatomic molecule **AB** that has a σ bond. This is represented as **A–B**.

II. The electronegativity of **A** is less than that of **B** ($X_A < X_B$).
X = electronegativity of the atom

III. The inter-nuclear distance between **A** and **B** atoms (d_{A-B}) of the **AB** molecule is given by the following equation.

$d_{A-B} = r_A + r_B - c(X_B - X_A)$

r = atomic radius, c = 9 pm

Note: d and r are measured in picometres (pm). (1 pm = 10^{-12} m)

Based on the above information, answer the following questions.

(i) What is the name used to identify the type of σ bond between **A** and **B**?

Polar covalent bond **(03)**

(ii) Show how fractional charges ($\delta+$ and $\delta-$) are located in the molecule **AB**.

$A^{\delta+}$— $B^{\delta-}$ **(03)**

(iii) Write the equation to calculate the dipole moment (μ) of molecule **AB** and show its direction.

$\mu = d_{AB} \times \delta$, **OR** $\mu = qr$, $\overrightarrow{A—B}$ **OR** $\overrightarrow{+\,\,-}$ **(01 + 01)**

(iv) Calculate the percentage of ionic character of the H–F bond in the HF molecule using the data given below.

Inter-nuclear distance of H_2 (d_{H-H}) = 74 pm Electronegativity of F = 4.0
Inter-nuclear distance of F_2 (d_{F-F}) = 144 pm Dipole moment of HF = 6.0×10^{-30} C m
Electronegativity of H = 2.1 Charge of an electron = 1.6×10^{-19} C

$\mu = d_{HF} \times \delta$, $H^{\delta+}$— $F^{\delta-}$

$r_H = \frac{d_{H2}}{2} = \frac{74}{2} = 37$ pm (02)

$r_F = \frac{d_{F2}}{2} = \frac{144}{2} = 72$ pm (02)

Therefore, $d_{HF} = 37 + 72 - 9(4.0 - 2.1)$ (01)
$= 109 - 9 \times 1.9$
$= 91.9$ pm (02)

$\mu = d_{HF} \times \delta$, 6.0×10^{-30} C m $= \delta \times 91.9 \times 10^{-12}$ m (01)

$\delta = \frac{6.0 \times 10^{-30}}{91.9 \times 10^{-12}} = 0.65 \times 10^{-19}$ (02)

% Ionic character = $\frac{0.65 \times 10^{-19}}{1.6 \times 10^{-19}} \times 100$ (01)
= **40.6%** (01)

OR

$r_H = \frac{d_{H2}}{2} = \frac{74}{2} = 37$ pm (02)

$r_F = \frac{d_{F2}}{2} = \frac{144}{2} = 72$ pm (02)

Therefore, $d_{HF} = 37 + 72 - 9(4.0 - 2.1)$ (01)
$= 109 - 9 \times 1.9$
$= 91.9$ pm (02)

$\mu_{ionic} = 1.6 \times 10^{-19}$ C $\times 91.9 \times 10^{-12}$ m (03)
$= 147.04 \times 10^{-31}$ C m

% Ionic character = $\frac{6 \times 10^{-30}}{147.04 \times 10^{-31}} \times 100$ (01)
= **40.8%** (01

2. (*a*) **A, B, C** and **D** are chlorides of *p*-block elements. These elements have atomic numbers less than 20. A description of the products ($P_1 - P_9$) formed when **A** is reacted with a limited amount of water and **B, C** and **D** are reacted with excess water are given below.

Compound	Description of products	
A	P_1	a compound with a covalent network structure
	P_2	a strong monobasic acid
B	P_3	a gas that turns red litmus blue
	P_4	a compound with bleaching properties
C	P_5	a tribasic acid
	P_6	a strong monobasic acid
D	P_7	a gas that turns acidic $KMnO_4$ solution colourless
	P_8	a colloidal solid
	P_9	a strong monobasic acid

(i) Identify **A, B, C** and **D** (give the chemical formulae).

A: $SiCl_4$ B: NCl_3 C: PCl_5 D: SCl_2 **(04 x 4)**

(ii) Give balanced chemical equations for the reactions of **A, B, C** and **D** with water to give products P_1 to P_9.

$SiCl_4 + 2H_2O \rightarrow SiO_2$ **(P_1)** $+ 4HCl$ **(P_2)** **(05)**

$NCl_3 + 3H_2O \rightarrow NH_3$ **(P_3)** $+ 3HOCl$ **(P_4)** **(05)**

$PCl_5 + 4H_2O \rightarrow H_3PO_4$ **(P_5)** $+ 5HCl$ **(P_6)** **(05)**

$2SCl_2 + 2H_2O \rightarrow SO_2$ **(P_7)** $+ S$ **(P_8)** $+ 4HCl$ **(P_9)** **(05)**

Note: Award marks if correct balanced equations are given.

(iii) Write balanced chemical equations for the following reactions.

I. P_1 with NaOH(aq)

$SiO_2 + 2NaOH \rightarrow Na_2SiO_3 + H_2O$ **(04)**

II. P_3 with Mg

$3Mg + 2NH_3 \rightarrow Mg_3N_2 + 3H_2$ **(04)**

III. P_7 with acidic $K_2Cr_2O_7$

$3(SO_2 + 2H_2O \rightarrow SO_4^{2-} + 4H^+ + 2e)$

$Cr_2O_7^{2-} + 14H^+ + 6e \rightarrow 2Cr^{3+} + 7H_2O$

$3SO_2 + Cr_2O_7^{2-} + 2H^+ \rightarrow 2Cr^{3+} + 3SO_4^{2-} + H_2O$ **(06)**

If only half reactions are given – part marks (02 + 02)

2(a): 50 marks

(*b*) A student is provided with bottles labelled **P, Q, R, S, T** and **U** containing aqueous solutions of $Al_2(SO_4)_3$, H_2SO_4, $Na_2S_2O_3$, $BaCl_2$, $Pb(Ac)_2$ and KOH **(not in order)**. Some useful observations for their identification on mixing two solutions at a time are given below. (Ac - Acetate ion)

	Solutions mixed	Observations
I	**T + R**	a clear colourless solution
II	**P + R**	a white precipitate
III	**T + S**	a gelatinous white precipitate
IV	**U + R**	a white precipitate
V	**P + Q**	a white precipitate, turns black on heating
VI	**P + U**	a white precipitate, dissolves on heating

(i) Identify **P** to **U**.

P: $Pb(Ac)_2$ **Q:** $Na_2S_2O_3$ **R:** H_2SO_4

S: $Al_2(SO_4)_3$ **OR** KOH **T:** KOH **OR** $Al_2(SO_4)_3$ **U:** $BaCl_2$

(05 X 6 = 30)

(ii) Give balanced chemical equations for each of the reactions **I** to **VI**.

I. $2KOH + H_2SO_4 \rightarrow K_2SO_4 + 2H_2O$ **OR** **(03)**
$Al_2(SO_4)_3 + H_2SO_4 \rightarrow$ No reaction

II. $Pb(Ac)_2 + H_2SO_4 \rightarrow PbSO_4\downarrow + 2HAc$ **(03)**

III. $6KOH + Al_2(SO_4)_3 \rightarrow 2Al(OH)_3\downarrow + 3K_2SO_4$ **(03)**

IV. $BaCl_2 + H_2SO_4 \rightarrow BaSO_4\downarrow + 2HCl$ **(03)**

V. Formation of white ppt

$Pb(Ac)_2 + Na_2S_2O_3 \rightarrow PbS_2O_3\downarrow + 2NaAc$ **(03)**

Turning black on heating

$PbS_2O_3 + H_2O \rightarrow PbS\downarrow + H_2SO_4$ **(02)**

VI. $Pb(Ac)_2 + BaCl_2 \rightarrow PbCl_2\downarrow + Ba(Ac)_2$ **(03)**

Note: Precipitates have to be shown by ↓ or as (s). If not, deduct (01) mark.

2(b): 50 marks

3. *(a)* A saturated aqueous solution of a sparingly soluble salt $AB_2(s)$ was prepared by stirring an excess amount of $AB_2(s)$ in 1.0 dm^3 of distilled water at 25 °C. The amount of $A^{2+}(aq)$ ions present in this saturated aqueous solution was found to be 2.0×10^{-3} mol.

(i) Write the equilibrium related to the dissolution of $AB_2(s)$ in the above system at 25 °C.

$AB_2(s) \rightleftharpoons A^{2+}(aq) + 2B^-(aq)$ **(05)**

(ii) Write the expression for the equilibrium constant for the equilibrium written in (i) above at 25 °C.

$K_{sp} = [A^{2+}(aq)][B^-(aq)]^2$ **(05)**

$K_C = \frac{[A^{2+}(aq)][B^-(aq)]^2}{[AB_2(s)]}$ **Note: If only K_c is given award (03 marks)**

(iii) Calculate the value of the equilibrium constant stated in (ii) above at 25 °C.

$[A^{2+}(aq)] = 2.0 \times 10^{-3}$mol dm^{-3} (04+01)

$[B^{-}(aq)] = 2[A^{2+}(aq)] = 4.0 \times 10^{-3}$mol dm^{-3} (04+01)

K_{sp} = 2.0×10^{-3}mol dm^{-3} × $(4.0 \times 10^{-3}$mol dm$^{-3})^2$ (05)

$\boldsymbol{K_{sp} = 3.2 \times 10^{-8}}$**mol^3 dm^{-9}** (05)

(iv) Another saturated aqueous solution of AB_2 was prepared by stirring an excess amount of $AB_2(s)$ in 2.0 dm^3 of distilled water at 25 °C. Giving reasons, predict the value of the equilibrium constant for this system.

K_{sp} = 3.2×10^{-8} mol^3 dm^{-9} (05)

K_{sp} is a constant at constant temperature (05)

and does not depends on the volume (05)

(v) A small amount of the strong electrolyte NaB(s) is added to a saturated aqueous solution of AB_2 at 25 °C. Giving reasons, predict whether the concentration of A^{2+}(aq) is increased or decreased.

Common-ion B^-(aq) is added (05)

∴ More AB_2(s) is formed to keep the K_{sp} constant **or** reverse reaction takes place (05)

$\boldsymbol{[A^{2+}(aq)], decreases}$ (05)

3(a): 60 marks

(b) In an aqueous solution, propanoic acid (C_2H_5COOH) ionizes as given below.

$C_2H_5COOH(aq) + H_2O(l) \rightleftharpoons C_2H_5COO^-(aq) + H_3O^+(aq)$

At 25 °C, K_a (propanoic acid) = 1.0×10^{-5}

(i) Write the expression for the equilibrium constant for the above reaction at 25 °C.

$$K_a = \frac{[C_2H_5COO^-(aq)][H_3O^+(aq)]}{[C_2H_5COOH\ \ (aq)]}$$ (05)

(ii) 100.0 cm^3 of an aqueous solution of C_2H_5COOH(aq) was prepared by dissolving 0.74 cm^3 of C_2H_5COOH in distilled water at 25 °C. Calculate the pH of the solution at 25 °C.
(C = 12; O = 16; H = 1; consider the density of C_2H_5COOH as 1.0 g cm^{-3})

mass of C_2H_5COOH(aq) = 0.74 cm^3 × 1.00 g cm^{-3} = 0.74 g

moles of C_2H_5COOH(aq) in 100 cm^3 = 0.74 g/ 74 g mol^{-1} = 0.01 mol (05)

∴ [C_2H_5COOH(aq)] = 0.10 mol dm^{-3} (05)

Consider the equilibrium:

$C_2H_5COOH(aq) + H_2O(l) \rightleftharpoons C_2H_5COO^-(aq) + H_3O^+(aq)$

Initial	*0.10*	*0*	*0*	*mol dm^{-3}*
Change	*-x*	*x*	*x*	*mol dm^{-3}*
At eqm	*0.10-x*	*x*	*x*	*mol dm^{-3}*

(05)

$$K_a = \frac{[C_2H_5COO^-(aq)][H_3O^+(aq)]}{[C_2H_5COOH\ \ (aq)]} = \frac{x.x}{0.10-x} = 1.0 \times 10^{-5}$$ (02)

$\frac{x^2}{0.10} = 1.0 \times 10^{-5}$ $(0.10 - x \sim 0.1)$ (03)

$x^2 = 1.0 \times 10^{-6}$

$x = 1.0 \times 10^{-3}\ mol\ dm^{-3} = H_3O^+(aq)]$ (05)

$pH = -log\ [H_3O^+(aq)] = 1.0 \times 10^{-3}$ (05)

$\boldsymbol{pH = 3.0}$ (05)

Note : Students may take –log of both sides of $K_a = \frac{[C_2H_5COO^-(aq)][H_3O^+(aq)]}{[C_2H_5COOH\ \ (aq)]}$ and calculate pH.

Award marks appropriately.

3(b): 40 marks

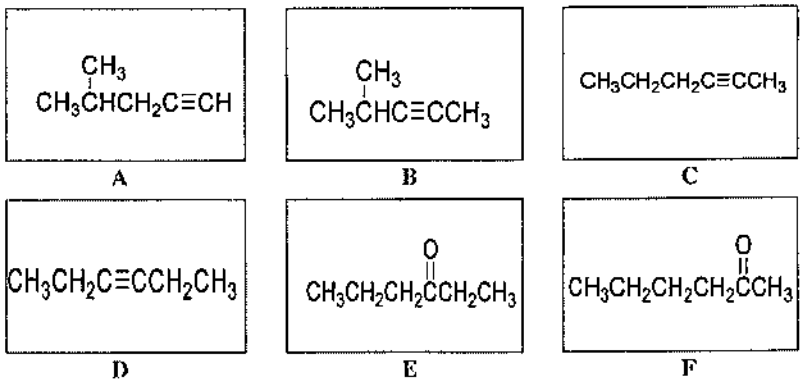
4. (*a*) **A, B, C** and **D** are structural isomers having the molecular formula C_6H_{10}. None of them show optical isomerism. All four isomers, **A, B, C** and **D** when treated with $HgSO_4$/dil. H_2SO_4 give products which react with 2,4-dinitrophenylhydrazine (2,4-DNP) to give coloured precipitates.

Only **A** gives a precipitate with ammonical $AgNO_3$. **A** has only one position isomer, which is **B**.

B is a chain isomer of **C**. **C** reacts with $HgSO_4$/dil. H_2SO_4 to give two products **E** and **F**.

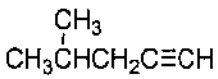
D reacts with $HgSO_4$/dil. H_2SO_4 to give only one product, which is **E**.

(i) Draw the structures of **A, B, C, D, E** and **F** in the boxes given below.



(06 x 6 = 36)

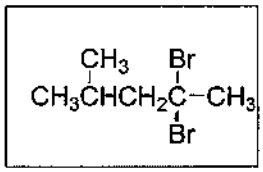
(ii) Which of the compounds **A, B, C** and **D** gives a product that does not show diastereoisomerism when reacted separately with H_2 / Pd-$BaSO_4$ / quinoline?



(05)

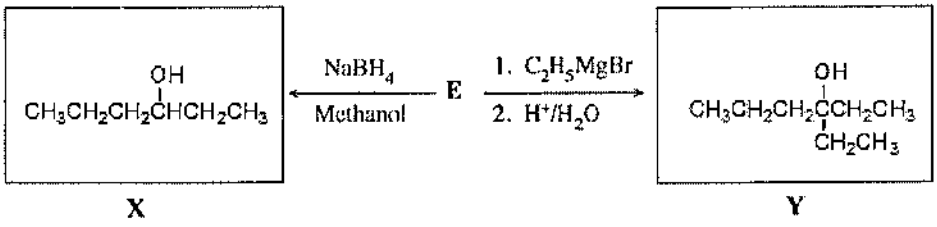
or Appropriate letter (**A**, **B**, **C** or **D**) identifying the correct structure

(iii) Draw, in the box given below, the structure of the product **G** obtained when **A** is reacted with excess HBr.



(05)

(iv) Draw the structures of products **X** and **Y** obtained in the following reactions of **E**, in the appropriate boxes.



Name a test to distinguish between **X** and **Y**.

(05 x 2 = 10)

Lucas test or

anh. $ZnCl_2$ / conc. HCl or

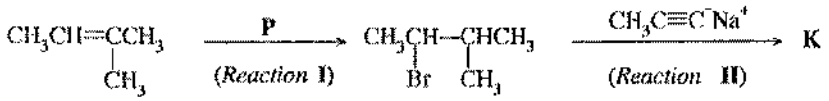
$H^+/K_2Cr_2O_7$ or

$H^+/KMnO_4$

(04)

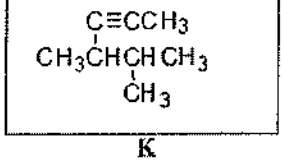
4(a): 60 marks

Sequence 1:



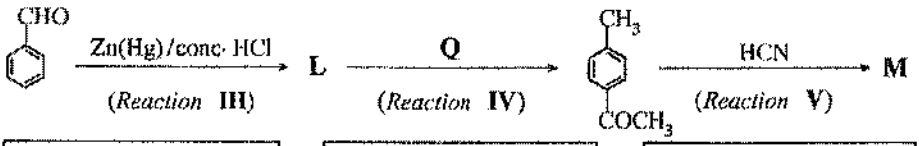
HBr/Peroxides

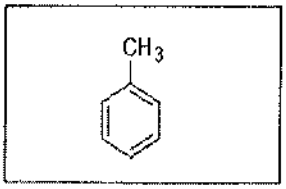
P



K

Sequence 2:

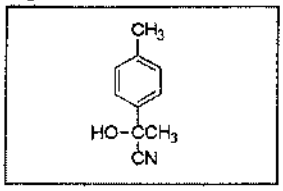




L

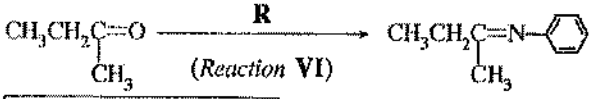
$CH_3COCl/AlCl_3$

Q



M

Sequence 3:



$C_6H_5NH_2$

R

Compounds/reagents **(05 x 6 = 30)**

(ii) Selecting from the reactions **I – VI**, give **one (01)** example for each of the following types of reactions.

Nucleophilic addition ..... Reaction V .....

Nucleophilic substitution ..... Reaction II .....

Reactions **(05 x 2 = 10)**

4(b): 40 marks

PART B — ESSAY

5. (*a*) A compound $XY_2Z_2(g)$ undergoes dissociation when heated to temperatures above 300 K as given below.

$$XY_2Z_2(g) \overset{\Delta}{\rightleftharpoons} XY_2(g) + Z_2(g)$$

A sample of 7.5 g of $XY_2Z_2(g)$ was placed in an evacuated 1.00 dm^3 rigid-closed container and the temperature was raised to 480 K.
Molar mass of $XY_2Z_2(g)$ is 150 g mol^{-1}. Use the approximate value of 4000 J mol^{-1} for *RT* at 480 K. Assume ideal gas behaviour for all gases.

(i) Calculate the number of moles of $XY_2Z_2(g)$ in the container before dissociation.

$7.5\ g/150\ g\ mol^{-1} = 5.0 \times 10^{-2}\ mol$ **(05)**

5(a) (i): 05 marks

(ii) When the above system reaches equilibrium at 480 K, the total number of moles in the container was found to be 7.5×10^{-2} mol. Calculate the number of moles of $XY_2Z_2(g)$, $XY_2(g)$ and $Z_2(g)$ in the equilibrium mixture at 480 K.

$XY_2Z_2(g)$	$\rightleftharpoons$	$XY_2(g)$	+	$Z_2(g)$		
Initial	*0.05*		*0*	*0*	*mol dm^{-3}*	**(05)**
Change	*-x*		*x*	*x*	*mol dm^{-3}*	
At eqm	*0.05-x*		*x*	*x*	*mol dm^{-3}*	**(05)**

Total number of moles $= 0.05+x = 7.5 \times 10^{-2}\ mol$ **(05)**

$x = 2.5 \times 10^{-2}\ mol$ **(05)**

$XY_2(g) = Z_2(g) = 2.5 \times 10^{-2}\ mol$ **(05)**

$XY_2Z_2(g) = 5.0 \times 10^{-2}\ mol - 2.5 \times 10^{-2}\ mol = 2.5 \times 10^{-2}\ mol$ **(05)**

5(a) (ii): 30 marks

(iii) Calculate the equilibrium constant K_c for the above reaction at 480 K.

$K_c = \frac{[XY_2(g)\ ][Z_2(g)]}{[XY_2Z_2(g)]}$ **(05)**

Concentration = $2.5 \times 10^{-2}\ mol\ dm^{-3}$ **(05)**

$K_c = \frac{2.5 \times 10^{-2}\ mol\ dm^{-3} \times 2.5 \times 10^{-2}\ mol\ dm^{-3}}{2.5 \times 10^{-2}\ mol\ dm^{-3}}$ **(05)**

$\boldsymbol{K_c = 2.5 \times 10^{-2}\ mol\ dm^{-3}}$ **(05)**

5(a) (iii): 20 marks

(iv) Calculate K_p for the equilibrium at 480 K.

$K_p = K_c\ (RT)^{\Delta n}$ **(05)**

$\Delta n = 1$ **(05)**

$K_p = 2.5 \times 10^{-2}\ mol\ dm^{-3} \times 4 \times 10^{3}\ J\ mol^{-1}$ **(05)**

$\boldsymbol{K_p = 1.0 \times 10^{5}\ Pa}$ **(05)**

iv. Alternative:

Total number of moles at equilibrium $= 7.5 \times 10^{-2}$ mol

$P_{Total} = (7.5 \times 10^{-2} \text{ mol} \times 4 \times 10^{3} \text{ J mol}^{-1}) / 1.0 \times 10^{-3} \text{ m}^3) = 3.0 \times 10^{5}$ Pa)

Number of moles of $XY_2Z_2(g) = XY_2(g) = Z_2(g) = 2.5 \times 10^{-2}$ mol

Mole fractions of $XY_2Z_2(g) = XY_2(g) = Z_2(g) = 1/3$

$P_i = X_i P_{total}$

$P_{XY2\,Z2(g)} = P_{XY2(g)} = P_{Z2(g)} = 1.0 \times 10^5$ Pa

$K_p = [P_{XY2(g)} = P_{Z2(g)}] / P_{XY2\,Z2(g)} = \mathbf{1.0 \times 10^5\ Pa}$

5(a): 75 marks

(*b*) For the reaction $XY_2Z_2(g) \rightarrow XY_2(g) + Z_2(g)$ described in (a), Gibbs free energies (*G*) at 480 K for $XY_2Z_2(g)$, $XY_2(g)$ and $Z_2(g)$ are –60 kJ mol^{-1}, –76 kJ mol^{-1} and –30 kJ mol^{-1}, respectively.

(i) Calculate ΔG (in kJ mol^{-1}) for the reaction at 480 K.

$$XY_2Z_2(g) \rightarrow XY_2(g) + Z_2(g)$$

$\Delta G_{rxn} = G_{products} - G_{reactants}$ **(05)**

$= (-76 + (-30)) - (-60) = \mathbf{-46\ kJ\ mol^{-1}}$ **(04+01)**

Note: No marks if ΔG^0_{rxn} is written.

5(b) (i): 10 marks

(ii) The magnitude of ΔS of the above reaction is 150 J K^{-1} mol^{-1} at 480 K. Calculate ΔH for the reaction at 480 K by using the appropriate sign (– or +) of ΔS.

ΔS must be positive (number of gaseous moles is higher in products) **(05)**

5(b) (ii): 05 marks

(iii) By using the sign (– or +) of ΔH obtained in (ii), explain whether this reaction is exothermic or endothermic.

$\Delta G = \Delta H - T\Delta S$ **(05)**

$-46\ kJ\ mol^{-1} = \Delta H - 480\ K \times 150 \times 10^{-3}\ kJ\ K^{-1}\ mol^{-1}$

$\Delta H = -46\ kJ\ mol^{-1} + 72\ kJ\ mol^{-1}$ **(04 +01)**

$\mathbf{\Delta H = +26\ kJ\ mol^{-1}}$ **(04+01)**

5(b) (iii): 15 marks

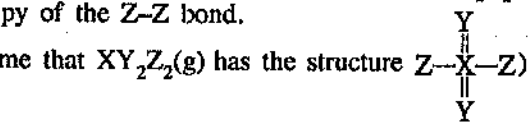
(iv) Deduce the enthalpy difference for the formation of $XY_2Z_2(g)$ from $XY_2(g)$ and $Z_2(g)$ at 480 K.

The reaction is endothermic **(05)**

because ΔH is positive **(05)**

5(b) (iv): 10 marks

(v) If the bond enthalpy of the X–Z bond in $XY_2Z_2(g)$ is +250 kJ mol^{-1}, calculate the bond enthalpy of the Z–Z bond.

(Assume that $XY_2Z_2(g)$ has the structure )

$$\Delta H = -26\ kJ\ mol^{-1}$$ **(09 +01)**

5(b) (v): 10 marks

(vi) If liquid XY_2Z_2 is used instead of gaseous XY_2Z_2, giving reasons, explain whether the value of ΔH obtained for the reaction $XY_2Z_2(l) \rightarrow XY_2(g) + Z_2(g)$ is equal to, or higher or lower than ΔH obtained in (ii).

$$\Delta H_{rxn} = \Delta H_{bonds\ formed} - \Delta H_{bonds\ broken}$$ **(05)**

$$\Delta H_{rxn} = \Delta H_{Z-Z} - 2\,\Delta H_{X-Z}$$ **(05)**

$$26\ kJ\ mol^{-1} = \Delta H_{Z-Z} - 2 \times 250\ kJ\ mol^{-1}$$

$$\Delta H_{Z-Z} = 526\ kJ\ mol^{-1}$$ **(04+01)**

(OR students may solve through an appropriate thermo cycle)

Higher **(05)**

It is necessary to supply energy to convert liquid to gas first **(05)**

(**or** $XY_2Z_2(l) \rightarrow XY_2Z_2(g)$ needs an extra energy)

5(b) (vi): 25 marks

5(b): 75 marks

6. (*a*) Consider the reaction given below occurring in a closed container at a given temperature *T*.

$$2N_2O_5(g) \rightarrow 4NO_2(g) + O_2(g)$$

(i) Write **three** expressions for the rate of reaction relevant to each of the compounds appearing in the reaction.

$$Rate = -\frac{\Delta[N_2O_5(g)]}{2\,\Delta t} = \frac{\Delta[NO_2(g)]}{4\,\Delta t} = \frac{\Delta[O_2(g)]}{\Delta t}$$ **(05)**

6(a) (i): 05 marks

(ii) This reaction was carried out at temperature *T* with an initial concentration of 0.10 mol dm^{-3} of $N_2O_5(g)$. It was found that 40% of the initial amount was decomposed after a period of 400 *s*.

I. Calculate the average rate of decomposition of $N_2O_5(g)$ in this time interval.

Decomposed amount = 0.10 mol dm^{-3} ×40/100 = 4.0×10^{-2} mol dm^{-3} **(05)**

Remaining concentration after 400 s = 6.0×10^{-2} mol dm^{-3} **(05)**

$$Average\ Rate = \frac{-(0.06 - 0.10)\ mol\ dm^{-3}}{(400-0)s} = 1.0 \times 10^{-4}\ mol\ dm^{-3}\ s^{-1}$$ **(05)**

II. Calculate average rates of formation of $NO_2(g)$ and $O_2(g)$.

$$\frac{\Delta[N_2O_5(g)]}{2\,\Delta t} = \frac{\Delta[NO_2(g)]}{4\,\Delta t}$$

$$\frac{\Delta[NO_2(g)]}{4\,\Delta t} = 2.0 \times 10^{-4}\ mol\ dm^{-3}\ s^{-1} \quad \textbf{(02)}$$

$$\frac{\Delta[O_2(g)]}{\Delta t} = \frac{\Delta[N_2O_5(g)]}{2\,\Delta t} = 5.0 \times 10^{-5}\ mol\ dm^{-3}\ s^{-1} \quad \textbf{(03)}$$

6(a) (ii): 20 marks

(iii) In another experiment, initial rates were measured for this reaction at 300 K and the results are given below.

$[N_2O_5(g)]$ / mol dm^{-3}	0.01	0.02	0.03
Initial rate / mol dm^{-3} s^{-1}	6.930×10^{-5}	1.386×10^{-4}	2.079×10^{-4}

Derive the rate law for the reaction at 300 K.

When the concentration were increased two and three times, rate increased two and three times, respectively. **(05)**

∴ Reaction is first order **(05)**

∴ Rate law : Rate = $k\,[N_2O_5(g)]$ **(05)**

*(**OR** $R_1/R_2 = 1/2$ :::::reaction is first order)*

6(a) (iii): 15 marks

(iv) Another experiment was carried out at 300 K with an initial concentration of 0.64 mol dm^{-3} of $N_2O_5(g)$. It was found that the concentration of $N_2O_5(g)$ which remained after a period of 500 s was 2.0×10^{-2} mol dm^{-3}.

I. Calculate the half-life ($t_{1/2}$) of the reaction at 300 K.

Order of concentration change = $0.64/2.0 \times 10^{-2} = 32 = (2)^5$ **(05)**

∴ Fraction of initial $N_2O_5(g) = (1/2)^5$ **(05)**

That is, 5 half-lives passed to reach this concentration **(05)**

∴ **$t_{1/2}$ = 500 s/5 = 100 s** **(05)**

II. Calculate the rate constant of the reaction at 300 K.

from iii,

Rate = $k\,[N_2O_5(g)] = 6.93 \times 10^{-5}\ mol\ dm^{-3}\ s^{-1} = k\ 0.01\ mol\ dm^{-3}$ **(05)**

***k* = 6.93 ×10^{-3} s^{-1}** **(04+01)**

OR

Reaction is first order

For first order reaction: $t_{1/2} = 0.693/k$ **(05)**

∴ $k = 0.693/100\ s = 6.93 \times 10^{-3}\ s^{-1}$ **(05)**

6(a) (iv): 30 marks

(v) This reaction proceeds through a mechanism involving the following elementary steps.

Step 1 : $N_2O_5(g) \rightleftharpoons NO_3(g) + NO_2(g)$: Fast

Step 2 : $NO_3(g) + NO_2(g) \rightarrow 2NO_2(g) + O(g)$: Slow

Step 3 : $N_2O_5(g) + O(g) \rightarrow 2NO_2(g) + O_2(g)$: Fast

Show that the above mechanism is consistent with the rate law of the reaction. **(8.0 *marks*)**

Step 1: $N_2O_5(g) \rightleftharpoons NO_3(g) + NO_2(g)$; *fast*

Step 2: $NO_3(g) + NO_2(g) \rightarrow 2NO_2(g) + O(g)$; *slow*

Step 3: $N_2O_5(g) + O(g) \rightarrow 2NO_2(g) + O_2(g)$; *fast*

From step 2 (Slow-step);

Rate= $k[NO_3(g)]\,[NO_2(g)]$ **(05)**

For step 1 (equilibrium)

$K_{eq} = \{[NO_3(g)]\,[NO_2(g)]\} / [N_2O_5(g)]$ **(05)**

We get, $K_{eq}[N_2O_5(g)] = \{[NO_3(g)]\,[NO_2(g)]\}$

$\therefore$ Rate= $k K_{eq}[N_2O_5(g)] = k'[N_2O_5(g)]$ **(05)**

This is a first order reaction which follows the rate low derived **(05)**

6(a) (v): 20 marks

6(a): 90 marks

(*b*) An ideal binary-liquid mixture was prepared by mixing two liquids of **A** and **B** in a closed evacuated container at temperature T. After establishing the equilibrium at temperature T, partial pressures of **A** and **B** in the vapour phase are P_A and P_B, respectively. At temperature T, the saturated vapour pressures of **A** and **B** are P_A^o and P_B^o, respectively. Mole fractions of **A** and **B** in solution are X_A and X_B, respectively.

(i) Show that $P_A = P_A^o X_A$

(Consider that the rates of vaporization and condensation are equal at equilibrium.)

Consider the above described vapor – liquid equilibrium of an ideal solution with components **A** and **B**. As the rate of evaporation equals the rate of condensation, we can write:

$$A_{(l)} \underset{r_c}{\overset{r_v}{\rightleftharpoons}} A_{(g)} \ldots\ldots (1)$$ **(05)**

r_v *and* r_c are the rates of vaporization and condensation, respectively of the component A. Considering (1), we can write;

$$r_v = k\left[A_{(l)}\right] = k_1 X_A$$ **(05)**

X_A is the mole fraction of A in solution

Likewise,

$$r_v' = k'[A_{(g)}] = k_2 P_A \quad \textbf{(05)}$$

P_A is the partial pressure of A in vapor phase.

At equilibrium

$$,r_v = r_v'$$

$$k_2 P_A = k_1 X_A \quad \textbf{(05)}$$

$$\therefore P_A = \frac{k_1}{k_2} X_A \text{ or } \therefore P_A = k X_A \quad \textbf{(05)}$$

when $X_A = 1$, $P_A = P_A^0$ = saturated vapor pressure of A

$$\therefore k = P_A^0 \quad \textbf{(05)}$$

$$\therefore P_A = P_A^0 X_A \quad \textbf{(05)}$$

6(b) (i): 35 marks

(ii) In the above system at 300 K, the total pressure was 5.0×10^4 Pa. The saturated vapour pressures of pure A and B at 300 K, are 7.0×10^4 Pa and 3.0×10^4 Pa, respectively.

I. Calculate the mole fraction of A in the liquid phase of the equilibrium mixture.

II. Calculate the vapour pressure of A in the equilibrium mixture.

(I) $P_{total} = P_A + P_B$ **(05)**

$$= X_A P^0_A + X_B P^0_B = X_A P^0_A + (1 - X_B) P^0_B \quad \textbf{(05)}$$

$$\therefore X_A = \frac{P_{total} - P_B^0}{P_A^0 - P_B^0} \quad \textbf{(05)}$$

$$= \frac{5 \times 10^4 - 3 \times 10^4}{7 \times 10^4 - 3 \times 10^4} = \frac{1}{2} \quad \textbf{(05)}$$

(II) $\therefore \quad P_A = P_A^0 X_A = \frac{1}{2} \times 7 \times 10^4 Pa = 3.5 \times 10^4 Pa$ **(05)**

6(b) (ii): 25 marks

6(b): 60 marks

7. (a) (i) To compare the properties of Electrolytic and Galvanic cells, copy and complete the following table using the given terms.

Terms: anode, cathode, positive, negative, spontaneous, non-spontaneous.

		Electrolytic cell	Galvanic cell
A.	Oxidation half reaction takes place at	Anode	Anode
B.	Reduction half reaction takes place at	Cathode	cathode
C.	Sign of E^0_{cell}	-ve	+ve
D.	Electron flow	**From anode to cathode**	**From anode to cathode**
E.	Spontaneity of reaction	Non-spontaneous	spontaneous

(2 ×10 = 20 marks)

7(a) (i): 20 marks

(ii) An electrochemical cell was constructed at 300 K by using a Zn(s) anode, an aqueous alkaline electrolyte and a porous Pt cathode which facilitates the collection of oxygen $O_2(g)$ from air as shown below. As the cell operates ZnO(s) is produced.

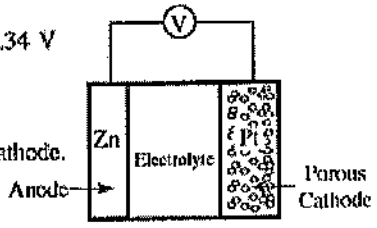
You are given that

$E^\circ_{ZnO(s)|Zn(s)|OH^-(aq)} = -1.31$ V and $E^\circ_{O_2(g)|OH^-(aq)} = +0.34$ V

Zn = 65 g mol^{-1}, O = 16 g mol^{-1} and

1 F = 96,500 C

I. Write the half-reactions occurring at anode and cathode.



anode : $Zn(s) + 2\ OH^-(aq) \rightarrow ZnO(s) + H_2O(l) + 2e$ **(05)**

Cathode ; $O_2(g) + 2\ H_2O(l) + 4e \rightarrow 4\ OH^-(aq)$ **(05)**

II. Write the overall cell reaction.

$2\ Zn(s) + O_2(g) \rightarrow 2ZnO(s)$ **(05)**

III. Calculate the cell potential E°_{cell} at 300 K.

$E^0_{cell} = E^0_R - E^0_L = E^0_{cathode} - E^0_{anode}$ **(05)**

$= 0.34\ V - (-1.31\ V) = 1.65V$ **(04+01)**

IV. State the direction of migration of $OH^-(aq)$ ions between the electrodes.

From anode to cathode (from Zn electrode to oxygen electrode) **(05)**

V. When the cell operates for a period of 800 s at 300 K, 2 mol of $O_2(g)$ are consumed.

A. Calculate the number of moles of electrons passing through the cell.

$$2 \text{ mol } O_2(g) \times \frac{4 \text{ mol ens}}{1 \text{ mol } O_2(g)} = 8 \text{ moles of electrons} \quad \textbf{(05)}$$

B. Calculate the mass of ZnO(s) formed.

$$\text{Mass of ZnO(s)} = \frac{8 \text{ mol ens} \times 96500 \text{ C}}{1 \text{ mol e} \times 800 \text{ s}} \times \frac{1 \text{ mol e}}{96500 \text{ C}} \times \frac{2 \text{ mol ZnO(s)}}{4 \text{ mol en}} \times \frac{81 \text{ g}}{1 \text{ mol ZnO}}$$

(05)

$= 324$ g **(04+01)**

OR

Mass of ZnO $= 4 \text{ mol} \times 81 \text{ g mol}$ **(05)**

$= 324$ g **(04+01)**

C. Calculate the current passing through the cell.

$I = q/t$ **(05)**

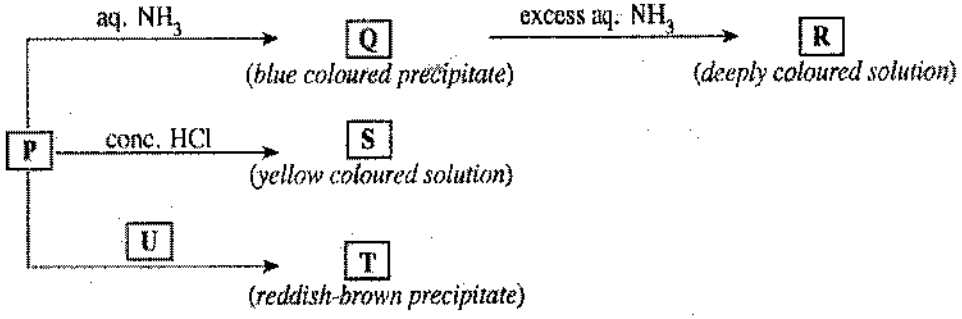
$$= \frac{8 \text{ mol ens} \times 96500 \text{ C}}{1 \text{ mol e} \times 800 \text{ s}} = 965 \text{ A}$$

(05)

7(a) (i): 55 marks

7(a): 75 marks

(*b*) A coloured complex ion **P** is formed when the salt $M(NO_3)_n$ is dissolved in distilled water. **M** is a transition element belonging to the 3*d* block. **P** undergoes the following reactions.



T and **U** are coordination compounds each containing four elements. **P**, **R** and **S** are complex ions.

(i) Identify the metal **M**. Give the oxidation state of **M** in complex ion **P**.

M = Cu **(10)**

Oxidation state: +2 **OR** Cu^{2+} **(03)**

7(b) (i): 13 marks

(ii) Give the value of n in $M(NO_3)_n$.

n = 2 **(03)**

7(b) (ii): 03 marks

(iii) Write the complete electronic configuration of **M** in complex ion **P**.

$1s^2 2s^2 2p^6 3s^2 3p^6 3d^9$ **(03)**

7(b) (iii): 03 marks

(iv) Write the chemical formulae of **P**, **Q**, **R**, **S**, **T** and **U**.

P: $[Cu(H_2O)_6]^{2+}$ **(04)**

Q: $Cu(OH)_2$ **(04)**

R: $[Cu(NH_3)_4]^{2+}$ **(04)**

S: $[CuCl_4]^{2-}$ **(04)**

T: $Cu_2[Fe(CN)_6]$

U: $K_4[Fe(CN)_6]$

7(b) (iv): 16 marks

(v) Give the IUPAC names of **P**, **R**, **S**, **T** and **U**.

P: hexaaquacopper(II) ion **(03)**

R: tetraamminecopper(II) ion **(03)**

S: tetrachloridocuprate(II) ion **(03)**

T: copper hexacyanoferrate(II)

U: potassium hexacyanoferrate(II)

7(b) (v): 12 marks

(vi) What is the colour of **P**?

pale blue **(04)**

7(b) (vi): 04 marks

(vii) What would you expect to observe in I and II given below?

I. When H_2S gas is passed into an acidic solution containing **P** at room temperature

black precipitate **(06)**

II. When the mixture obtained in I above is heated with dilute HNO_3 after the removal of dissolved H_2S

pale blue solution **(04)**
solution is turbid/ pale yellow or milky/ white precipitate **(02)**
or
turbid pale blue solution **(06)**

7(b) (vii): 12 marks

(viii) Briefly describe a method with the aid of balanced chemical equations for determining the concentration of M^{n+} present in an aqueous solution, using the following chemicals.
KI, $Na_2S_2O_3$ and starch.

Add excess KI **(01)**
to an aqueous solution of volume V_1 cm^3 containing M^{n+} **(01)**
Here, $M^{n+} = Cu^{2+}$
Titrate the liberated I_2 **(01)**
with $Na_2S_2O_3$ whose concentration is known (M mol dm^{-3}) **(01)**
with starch as the indicator **(01)**

$2Cu^{2+} + 2I^- \rightarrow 2Cu^+ + I_2$ ----------(1) **(01)**

$I_2 + 2S_2O_3^{2-} \rightarrow S_4O_6^{2-} + 2I^-$ ----------(2) **(01)**

$2Cu^{2+} + 2S_2O_3^{2-} \rightarrow S_4O_6^{2-} + 2Cu^+$ ----------(3) **(02)**

OR

$2Cu^{2+} + 4I^- \rightarrow 2CuI\downarrow + I_2$ ----------(1a) **(01)**

$I_2 + 2S_2O_3^{2-} \rightarrow S_4O_6^{2-} + 2I^-$ ----------(2a) **(01)**

$2Cu^{2+} + 2S_2O_3^{2-} \rightarrow 2CuI\downarrow + S_4O_6^{2-}$ ----------(3a) **(02)**

Note: If correct overall equation is given, award the part marks for half equations as well.

From both (3) or (3a) $Cu^{2+} \equiv S_2O_3^{2-}$ **(01)**

Let the burette reading of $S_2O_3^{2-}$ be V_2 cm^3 **(01)**

Therefore, moles of $S_2O_3^{2-} = \frac{V_2}{1000} \times M$ **(01)**

Therefore, moles of $Cu^{2+} = \frac{V_2}{1000} \times M$ **(01)**

Therefore, $[Cu^{2+}] = \frac{V_2}{1000} \times M \times \frac{1000}{V_1}$ **(01)**

$= \frac{MV_2}{V_1}$ mol dm^{-3} **(01)**

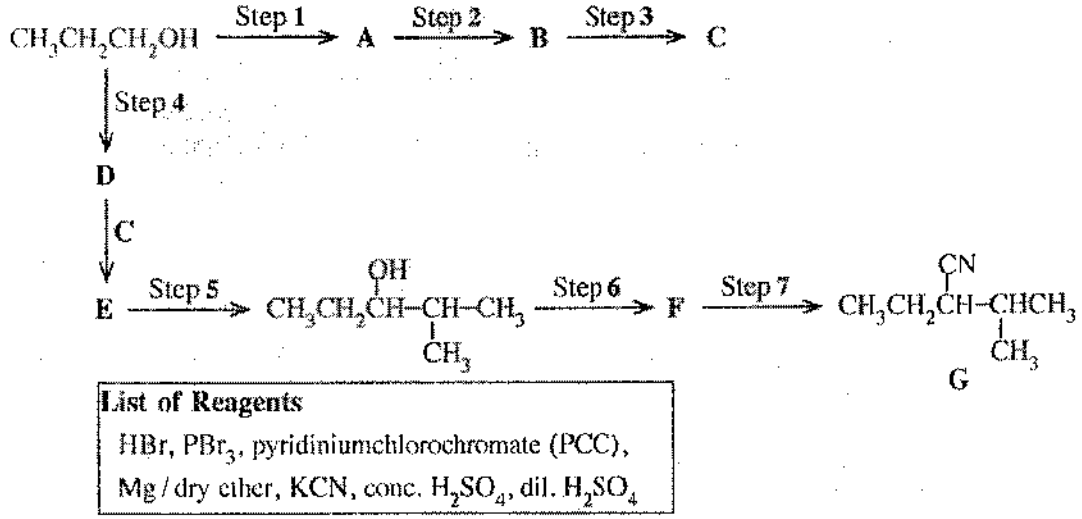
Note: The above explanation could be given in words.

7(b)(viii) : 15 marks

7(b): 75 marks

8. (*a*) (i) Given below is a reaction scheme for the synthesis of compound **G** using $CH_3CH_2CH_2OH$ as the only organic starting compound.

Complete the reaction scheme by drawing the structures of compounds **A, B, C, D, E** and **F** and writing the appropriate reagents for steps 1 – 7, selected only from those given in the list.



List of Reagents

HBr, PBr_3, pyridiniumchlorochromate (PCC), Mg / dry ether, KCN, conc. H_2SO_4, dil. H_2SO_4

Compounds, A - F

A = $CH_3CH{=}CH_2$

B = $CH_3CH(CH_3)Br$

C = $CH_3CH(CH_3)MgBr$

D = CH_3CH_2CHO

E = $CH_3CH_2C(OMgBr)CH(CH_3)CH_3$

F = $CH_3CH_2C(Br)CH(CH_3)CH_3$

Reagents:

Step 1 = conc.H_2SO_4

Step 2 = HBr

Step 3 = Mg / dry ether

Step 4 = PCC

Step 5 = dil.H_2SO_4

Step 6 = PBr_3

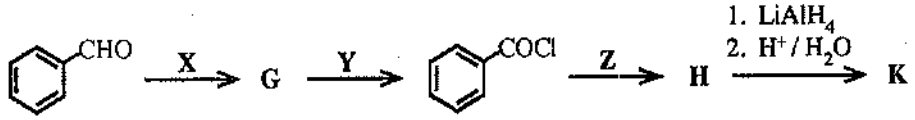
Step 7 = KCN

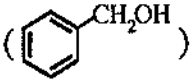
Compounds/Reagents **(04 x 13 = 52 marks)**

8(a) (i): 52 marks

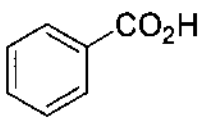
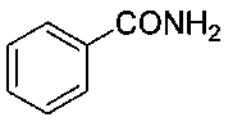
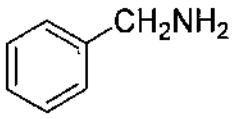
(ii) Consider the following series of reactions.

Draw the structures of compounds **G, H** and **K**. Give the reagents **X, Y** and **Z**.



Note that **K** gives benzyl alcohol () when reacted with $NaNO_2$/ dil. HCl.

Compounds **G, H** and **K**

G =  **H** =  **K** = 

Reagents

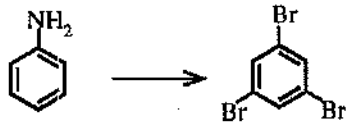
X = $H^+ / K_2Cr_2O_7$ / or $H^+ / KMnO_4$ or H^+/CrO_3 **Y** = PCl_5 or PCl_3 **Z** = NH_3

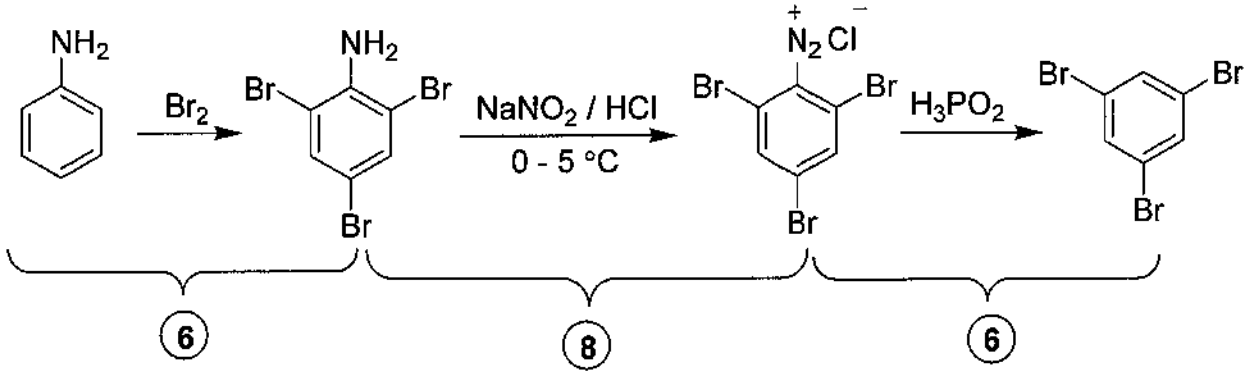
Compounds/Reagents **(04 x 6 = 24 marks)**

8(a) (ii): 24 marks

8(a): 76 marks

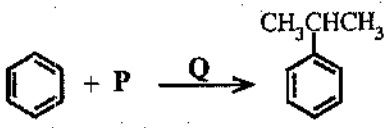
(*b*) (i) Show how the following conversion could be carried out in **not more than three** steps





8(b) (i) 20 marks

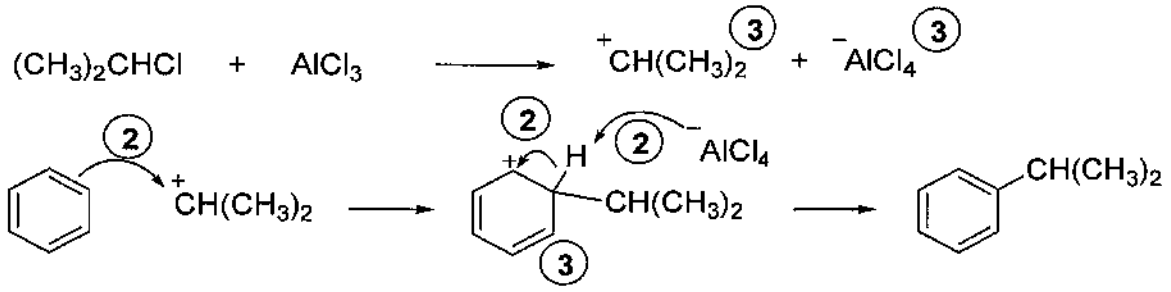
(ii) Consider the following reaction.



Identify the chemical substances **P** and **Q** necessary to carry out this reaction.
Write the mechanism of this reaction.

P + Q = $(CH_3)_2CHCl$ + $AlCl_3$ (5)

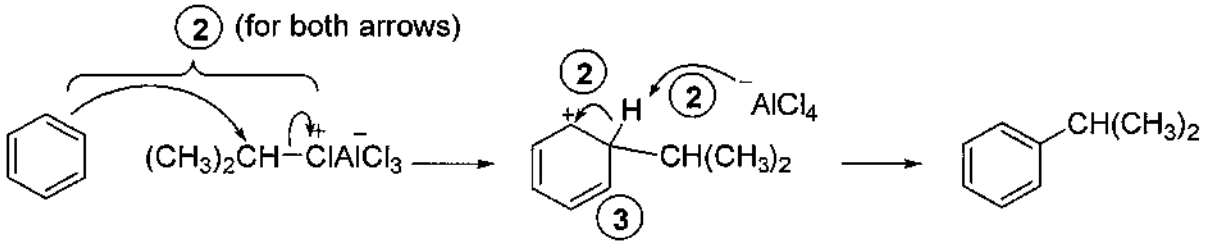
P + Q = **(05)**



Intermediates **03 x 3 = 09**
Arrows **02 x 3 = 06**

Alternative answer:

IF the student has written the electrophile as R–Cl molecule polarized by coordinating to $AlCl_3$, only the marks allocated for the last two steps may be awarded as given below.



02+02+02+03 = 09

The electrophile may be written as:

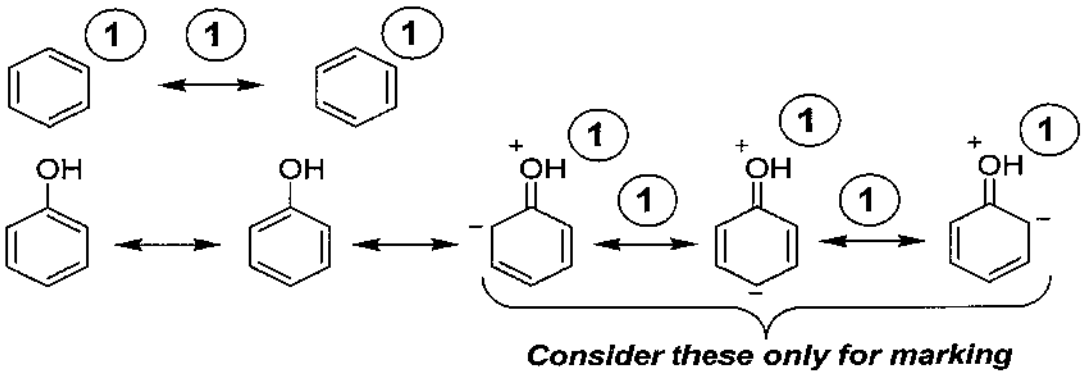
$(CH_3)_2CH-\overset{+}{C}l\overset{-}{A}lCl_3$ or $(CH_3)_2CH-\overset{+}{C}l-\overset{-}{A}lCl_3$ or $(CH_3)_2CH-\overset{\delta+}{C}l\cdots\overset{\delta-}{A}lCl_3$

8(b)(ii): 20 marks

8(b): 40 marks

(c) (i) Explain why phenol is more reactive in electrophilic substitution reactions than benzene, by considering their resonance hybrids.

Structures of benzene and phenol can be illustrated as follows.

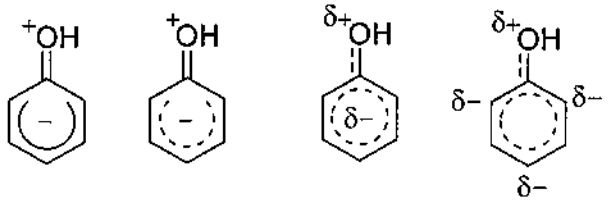


Consider these only for marking

Resonance structures and Double headed arrows **01** x 8 = **08**

OR

Alternative to the resonance hybrid of benzene (3)



Any one structure (5)

Alternative to the resonance hybrid of phenol

The benzene ring of phenol is more reactive towards electrophiles than benzene itself because:

The benzene ring in phenol is electron rich compared to benzene due to the

Delocalization of lone pair of electrons on the oxygen atom

Over the benzene ring of phenol

04 x 3 = 12

8(c)(i): 20 marks

(ii) Illustrate the difference in reactivity between phenol and benzene as given in (i) above by means of a suitable reaction.

Phenol **reacts with bromine at room temperature/ decolorizes bromine / gives a white precipitate with bromine water**

Benzene **does not react with bromine at room temperature / does not decolorize bromine / does not give a white precipitate with bromine water**

OR

Benzene reacts with bromine **(only) in the presence of a Lewis catalyst**

Reaction of phenol with bromine occurs **even in the absence of Lewis catalyst**

OR

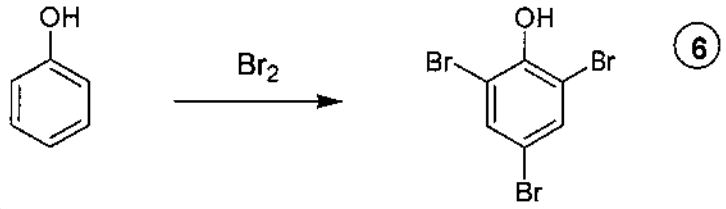
Nitration of phenol takes place at **room temperature /** at **20 °C / without heating with dilute HNO_3 (20% HNO_3).**

Benzene **does not react with dilute HNO_3**

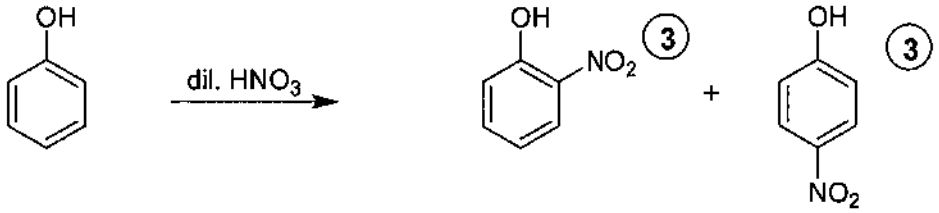
04 x 2 = 08

8(c): (ii) 08 marks

(iii) Draw the structure(s) of product(s) you described in the reaction in (ii) above.



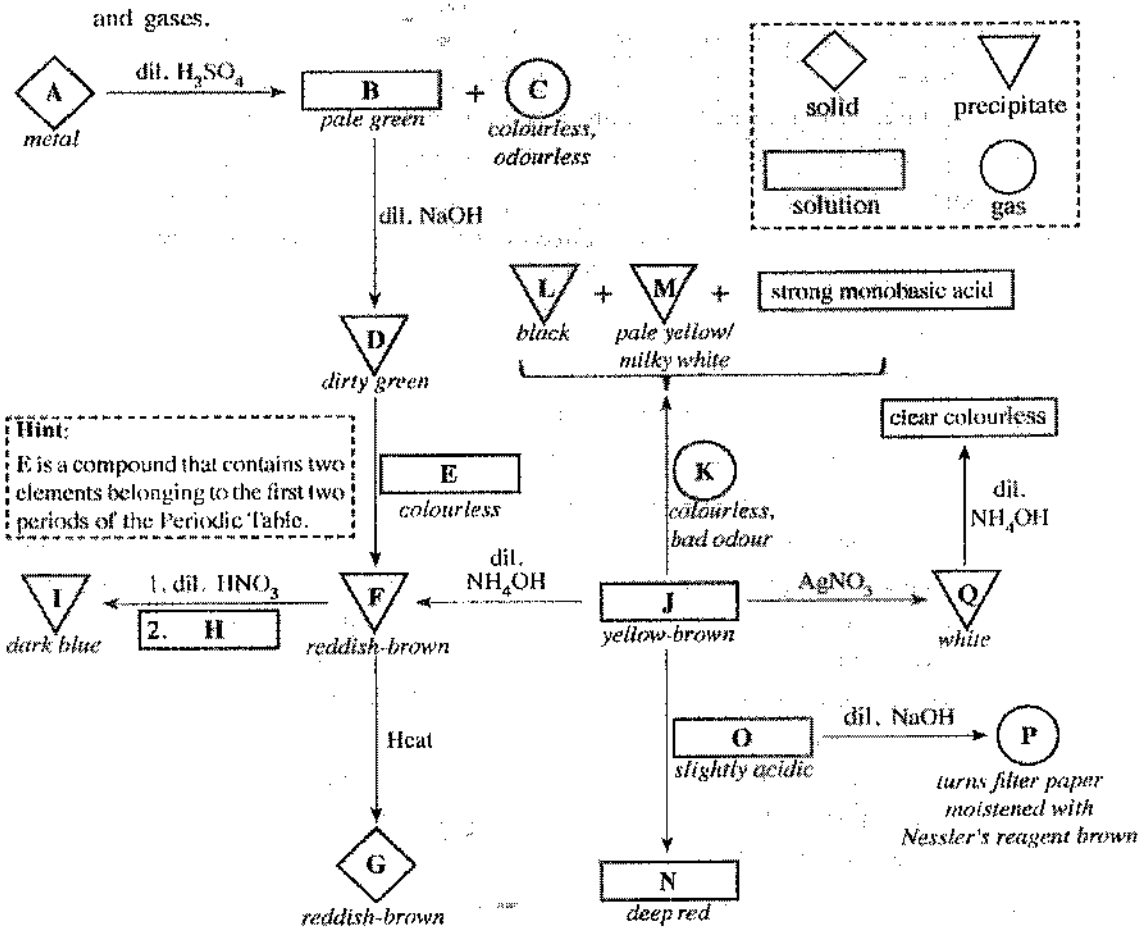
OR



8(c) (iii): 06 marks

8(c): 34 marks

9. (*a*) (i) Write the chemical formulae of the substances **A** – **Q** given in the flow chart below.
(**Note**: Chemical equations and reasons are not expected for the identification of substances **A** – **Q**.)
The symbols given in the box (dash lines) are used to represent solids, precipitates, solutions and gases.



A: Fe	**B:** $FeSO_4$ **or** $[Fe(H_2O)_6]SO_4$ **or** $[Fe(H_2O)_6]^{2+}$	**C:** H_2	**D:** $Fe(OH)_2$
E: H_2O_2	**F:** $Fe(OH)_3$	**G:** Fe_2O_3	**H:** $K_4[Fe(CN)_6]$
I: $Fe_4[Fe(CN)_6]_3$ **or** $KFe[Fe(CN)_6]$	**J:** $FeCl_3$	**K:** H_2S	**L:** FeS
M: S **or** S_8	**N:** $Fe(SCN)_3$ **or** $[Fe(SCN)(H_2O)_5]^{2+}$ **or** $[Fe(SCN)]^{2+}$	**O:** NH_4SCN	**P:** NH_3
Q: AgCl			

(04 marks x 17 = 68 marks)

9 (a) (i) : 68 marks

(ii) Write the complete electronic configuration of **A**.

$1s^2 2s^2 2p^6 3s^2 3p^6 3d^6 4s^2$ **(02)**

(iii) State the function of **E** in the conversion of **D** to **F**. Give the relevant balanced chemical equations for the stated function.

E: H_2O_2 function – oxidizing agent **(02)**

$2(Fe(OH)_2 + H_2O \rightarrow Fe(OH)_3 + H^+ + e)$

$H_2O_2 + 2H^+ + 2e \rightarrow 2H_2O$

$2Fe(OH)_2 + H_2O_2 \rightarrow 2Fe(OH)_3$ **(03)**

OR

$2(Fe^{2+} \rightarrow Fe^{3+} + e)$

$H_2O_2 + 2H^+ + 2e \rightarrow 2H_2O$

$2Fe^{2+} + H_2O_2 + 2H^+ \rightarrow 2Fe^{3+} + 2H_2O$ **(03)**

(Half reactions (01) each if written)

9 (a) (ii & iii) marks :07

9 (a) 75 marks

(*b*) The solid **X** contains only Cu_2S and CuS. The following procedure was used to determine the percentage of Cu_2S in **X**.

Procedure

A 1.00 g portion of solid **X** was treated with 100.00 cm^3 of 0.16 mol dm^{-3} $KMnO_4$ in dilute H_2SO_4 medium. This reaction gave Mn^{2+}, Cu^{2+} and SO_4^{2-} as products. Thereafter, the excess $KMnO_4$ in this solution was titrated with 0.15 mol dm^{-3} Fe^{2+} solution. The volume required for the titration was 35.00 cm^3.

(i) Write the balanced ionic equations for the reactions taking place in the above procedure.

Reaction of Cu_2S with MnO_4^-

$2Cu^+ \rightarrow 2Cu^{2+} + 2e$ ---------(1) **(03)**

$S^{2-} + 4H_2O \rightarrow SO_4^{2-} + 8H^+ + 8e$ ---------(2) **(03)**

OR

(1) + (2)

$2Cu^+ + S^{2-} + 4H_2O \rightarrow 2Cu^{2+} + SO_4^{2-} + 8H^+ + 10e$ ------(3)

$2(MnO_4^- + 8H^+ + 5e \rightarrow Mn^{2+} + 4H_2O)$ --------(4) **(03)**

(3) + (4)

$2Cu^+ + S^{2-} + 2MnO_4^- + 8H^+ \rightarrow 2Cu^{2+} + SO_4^{2-} + 2Mn^{2+} + 4H_2O$ **(05)**

OR

$Cu_2S + 2MnO_4^- + 8H^+ \rightarrow 2Cu^{2+} + SO_4^{2-} + 2Mn^{2+} + 4H_2O$

(If only this equation is written award the full 14 marks)

Reaction of CuS with MnO_4^-

$5(S^{2-} + 4H_2O \rightarrow SO_4^{2-} + 8H^+ + 8e)$ ---------(5)

$8(MnO_4^- + 8H^+ + 5e \rightarrow Mn^{2+} + 4H_2O)$ ---------(6)

(5)+ (6)

$5S^{2-} + 8MnO_4^- + 24H^+ \rightarrow 5SO_4^{2-} + 8Mn^{2+} + 12H_2O$ **(05)**

OR

$5CuS + 8MnO_4^- + 24H^+ \rightarrow 5CuSO_4 + 8Mn^{2+} + 12H_2O$

Reaction of Fe^{2+} with MnO_4

$5(Fe^{2+} \rightarrow Fe^{3+} + e)$ ----------(7) **(03)**

$MnO_4^- + 8H^+ + 5e \rightarrow Mn^{2+} + 4H_2O$ ----------(8)

(7) + (8)

$5Fe^{2+} + MnO_4^- + 8H^+ \rightarrow Mn^{2+} + 5Fe^{3+} + 4H_2O$ **(05)**

9 (b)(i) : 27 marks

OR

Reaction of Cu^{+} with MnO_4^-

$5(Cu^{+} \rightarrow Cu^{2+} + e)$ ----------(1a) **(03)**

$MnO_4^- + 8H^{+} + 5e \rightarrow Mn^{2+} + 4H_2O$ ----------(2a) **(03)**

(1a) + (2a)

$5Cu^{+} + MnO_4^- + 8H^{+} \rightarrow 5Cu^{2+} + Mn^{2+} + 4H_2O$ **(05)**

Reaction of S^{2-} with MnO_4^-

$5(S^{2-} + 4H_2O \rightarrow SO_4^{2-} + 8H^{+} + 8e)$ ----------(5) **(03)**

$8(MnO_4^- + 8H^{+} + 5e \rightarrow Mn^{2+} + 4H_2O)$ ----------(6)

(5)+ (6)

$5S^{2-} + 8MnO_4^- + 24H^{+} \rightarrow 5SO_4^{2-} + 8Mn^{2+} + 12H_2O$ **(05)**

Reaction of Fe^{2+} with MnO_4^-

$5(Fe^{2+} \rightarrow Fe^{3+} + e)$ ----------(7) **(03)**

$MnO_4^- + 8H^{+} + 5e \rightarrow Mn^{2+} + 4H_2O$ ----------(8)

(7) + (8)

$5Fe^{2+} + MnO_4^- + 8H^{+} \rightarrow Mn^{2+} + 5Fe^{3+} + 4H_2O$ **(05)**

Note: If only the overall reaction is written correctly, award the marks due to the half reactions as well.

b (b)(i): 27 marks

(ii) Based on the answers to (i) above, determine the molar ratio between,

I. Cu_2S and $KMnO_4$
II. CuS and $KMnO_4$
III. Fe^{2+} and $KMnO_4$

Molar ratios

$\frac{Cu_2S}{MnO_4^-} = \frac{1}{2}$ $\frac{CuS}{MnO_4^-} = \frac{5}{8}$ $\frac{Fe^{2+}}{MnO_4^-} = \frac{5}{1}$ **(05 x 3)**

OR

$Cu_2S : MnO_4^- = 1 : 2$ $CuS : MnO_4^- = 5 : 8$ $Fe^{2+} : MnO_4^- = 5 : 1$

9 (b)(ii): 15 marks

(iii) Calculate the percentage by weight of Cu_2S in **X**. (Cu = 63.5, S = 32)

Let the number of moles of Cu_2S and CuS be n_1 and n_2 respectively in the 1.0 g of sample **X**

Molar mass of Cu_2S = (2 x 63.5) + 32 = 159 **(02)**

Molar mass of CuS = 63.5 + 32 = 95.5 **(02)**

$159n_1 + 95.5n_2 = 1.0$ ----------(9) **(02)**

Moles of Fe^{2+} reacted $= \frac{0.15}{1000} \times 35.0$ **(02)**

Moles of MnO_4^- $= \frac{0.15}{1000} \times 35.0 \times \frac{1}{5}$ **(02)**

Moles of MnO_4^- reacted with Cu_2S and CuS

$= \frac{0.16}{1000} \times 100.0 - \frac{0.15}{1000} \times 35.0 \times \frac{1}{5}$ **(02)**

$= 0.016 - 0.001$ **(02)**

$= 0.015$ mol **(02)**

Based on molar ratios

$2n_1 + \frac{8}{5} n_2 = 0.015$ ----------(10) **(02)**

(9) + (10)

$2n_1 + \frac{8}{5} \frac{(1-159\,n1)}{95.5} = 0.015$ **(02)**

$2 \times 5 \times 95.5\, n_1 + 8(1-159n_1) = 0.015 \times 95.5 \times 5$ **(02)**

$955n_1 + 8 - 1272n_1 = 7.1625$

$317n_1 = 0.84$

$n_1 = 0.0027$ **(02)**

Weight of Cu_2S = 0.0027 x 159 g **(02)**

= 0.43 g **(02)**

% Cu_2S $= \frac{0.43}{1.0} \times 100$ **(02)**

= 43% **(03)**

9 (b)(iii) : 33 marks

OR

Moles of Fe^{2+} $= \frac{0.15}{1000} \times 35.0$ **(02)**

Moles of MnO_4^- remaining $= \frac{0.15}{1000} \times 35.0 \times \frac{1}{5}$ **(02)**

Moles of MnO_4^- added $= \frac{0.16}{1000} \times 100.0$ **(02)**

Moles of MnO_4^- reacted with Cu_2S and CuS

$= \frac{0.16}{1000} \times 100.0 - \frac{0.15}{1000} \times 35.0 \times \frac{1}{5}$ **(02)**

$= 0.016 - 0.001$ **(02)**

$= 0.015$ mol **(02)**

Consider the masses of Cu_2S and CuS to be p and q respectively.

$p + q = 1.0$ g ----------(9a) **(02)**

Molar mass of $Cu_2S = (2 \times 63.5) + 32 = 159$ **(02)**

Molar mass of $CuS = 63.5 + 32 = 95.5$ **(02)**

$\frac{2p}{159} + \frac{8q}{95.5 \times 5} = 0.015$ ----------(10a) **(02)**

From (9a) & (10a)

$\frac{2p}{159} + \frac{8(1-p)}{95.5 \times 5} = 0.015$ **(02)**

$2p \times 5 \times 95.5 + 8 \times 159(1-p) = 0.015 \times 5 \times 159 \times 95.5$ **(02)**

$955p - 1272p = 1138.84 - 1272$ **(02)**

$317p = 133.16$

$p = \frac{133.16}{317} = 0.42$ **(02)**

% $Cu_2S = \frac{0.42}{1000} \times 100.0$ **(02)**

= **42%** **(03)**

9 (b)(iii) : 33 marks

OR

Let the number of moles of Cu_2S and CuS be n_1 and n_2 respectively in the 1.0 g of **X**

$5Cu^+ + MnO_4^- + 8H^+ \rightarrow 5Cu^{2+} + Mn^{2+} + 4H_2O$

$5S^{2-} + 8MnO_4^- + 24H^+ \rightarrow 5SO_4^{2-} + 8Mn^{2+} + 12H_2O$

$5Fe^{2+} + MnO_4^- + 8H^+ \rightarrow Mn^{2+} + 5Fe^{3+} + 4H_2O$

Moles of MnO_4^- added $= \frac{0.16}{1000} \times 100.0 = 0.016$ **(02)**

Moles of Fe^{2+} reacted $= \frac{0.15}{1000} \times 35.0 = 0.005$ **(02)**

Moles of MnO_4^- remaining $= \frac{0.15}{1000} \times 35.0 \times \frac{1}{5} = 0.001$ **(02)**

Moles of MnO_4^- reacted $= 0.016 - 0.001 = 0.015$ **(02)**

Molar mass of $Cu_2S = (2 \times 63.5) + 32 = 159$ **(02)**

Molar mass of $CuS = 63.5 + 32 = 95.5$ **(02)**

$159n_1 + 95.5n_2 = 1$ ----------(1) **(02)**

Moles of $Cu^+ = 2n_1$

Therefore, moles of MnO_4^- reacted $= \frac{2n_1}{5}$

Moles of $S^{2-} = n_1 + n_2$ **(02)**

Therefore, moles of MnO_4^- reacted with $S^{2-} = \frac{8(n_1 + n_2)}{5}$

Therefore, total moles of MnO_4^- reacted $= \frac{10n_1 + 8n_2}{5}$ **(02)**

$\frac{(10n_1 + 8n_2)}{5}$ mol $= 0.015$ mol **(02)**

$10n_1 + 8n_2 = 0.075$ mol ----------(2) **(02)**

(1) x 8 – (2) x 95.5

$1272\ n_1 - 955\ n_1 = 8 - 7.14$ **(02)**

$317n_1 = 0.86$ Therefore, $n_1 = \frac{0.86}{317}$

Therefore, moles of Cu_2S in 1 g $= \frac{0.86}{317}$ **(02)**

Mass of Cu_2S $= 0.86 \times 159$ g **(02)**

% of Cu_2S $= \frac{0.86}{317} \times 159 \times 100\%$ **(02)**

$= \mathbf{43\%}$ **(03)**

9 (b)(iii): 33 marks

9(b): 75 marks

10. (*a*) The following questions are based on the properties of titanium dioxide (TiO_2) and its manufacture carried out by the "Chloride Process".

(i) Name the raw materials used in this process.

Rutile **(02)**

Coke **(02)**

Cl_2 **(02)**

O_2 **(02)**

10 (a) (i): 08 marks

(ii) Briefly describe the manufacturing process of TiO_2 giving balanced chemical equations where applicable.

<u>Chlorination</u>

Removal of water at 200 °C / 300 °C **(02)**

Heating of rutile and coke mixture at 900 °C / 950 °C **(02)**

$TiO_2(s) + C(s) \rightarrow Ti(s) + CO_2(g)$ --------(A) **(03)**

Stream of chlorine is passed over mixture of rutile and coke **(02)**

$Ti(s) + 2Cl_2(g) \rightarrow TiCl_4(g)$ --------(B) **(03)**

OR

Reactions (A) and (B) can be combined.

$TiO_2(s) + C(s) + 2Cl_2 \rightarrow TiCl_4(g) + CO_2(g)$ **(06)**

For three descriptions above **(02 x 3)**

After removal of dust particles, $TiCl_4$ gaseous mixture is cooled and liquid $TiCl_4$ is separated. **(02)**

Oxidation

$TiCl_4$ is reacted with oxygen and TiO_2 is regenerated.

$TiCl_4(g) + O_2(g) \rightarrow TiO_2(s) + 2Cl_2(g)$ **(03)**

Cl_2 is re-used in chlorination. **(02)**

10 (a) (ii): 19 marks

(iii) State **three** properties of TiO_2 and give one use each, relevant to each property.

- White colour – as a pigment in paint, plastic goods and paper, paper
- High refractive index – as a pigment
- Chemically inert – as a pigment in medicine and toothpaste
- Prevents the reach of UV rays to skin – produce substances to prevent sunburn

Any three properties **(02 x 3 = 06)**

One use for each property **(02 x 3 = 06)**

10 (a) (iii): 12 marks

(iv) If you were to consider establishing a TiO_2 manufacturing plant in Sri Lanka, state **three** requirements that need to be fulfilled.

- Availability of raw material
- Capitol
- Labour force
- Technology
- Storage conditions
- Minimize environmental pollution
- Transport facilities
- Waste product management

Any three **(02 x 3 = 06)**

10 (a) (iv): 06 marks

(v) Does the manufacturing process described in (ii) above contribute to global warming? Justify your answer.

Yes. **(02)**

CO_2 is produced and given out to the environment in the oxidation of coke **(03)**

10 (a) (v): 05 marks

10(a): 50 marks

(*b*) Currently, global warming due to change in greenhouse effect is significantly greater than that before the industrial revolution.

(i) Explain briefly what is meant by greenhouse effect.

Heating of earth (01) by infrared absorbing gases (01) in the atmosphere by trapping energy (IR radiation) (02) reradiated from the earth surface (02).

10 (b) (i): 06 marks

(ii) Identify the major environmental problem that occurs due to global warming.

Climate change **(03)**

10 (b) (ii): 03 marks

(iii) State **two main** natural gases that contribute to global warming.

CO_2, CH_4, and N_2O any two **(03 + 03)**

10 (b) (iii): 06 marks

(iv) Explain briefly how microorganisms contribute to the release of the gases you stated in (iii).

CO_2- Action of aerobic bacteria on organic substances/ plant materials/ and animal materials

CH_4- Action of anaerobic bacteria on organic substances/ materials

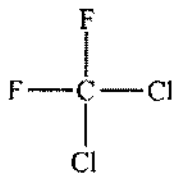
N_2O- Action denitrifying bacteria on ammonia/ nitrogen fertilizers(urea)/ and nitrogen containing substances.

Any two (04 + 04)

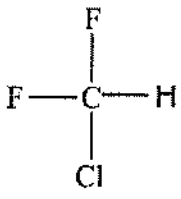
10 (b) (iv): 08 marks

(v) In addition to the gases you stated in (iii), name **two** classes of synthetic volatile compounds that directly contribute to the global warming, and selecting one compound from each class, draw their structures.

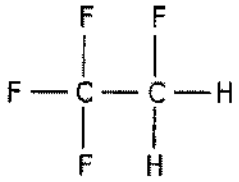
CFC, HFC, HCFC



CFC,



HFC,



HCFC

Any two (03 for class +03 for the structure)
(03 x 4 =12 marks)
No marks for the structure if the class is wrong

Note In addition to these compounds award marks for the following structures on each class.

CFC - Any saturated organic compound that contain one or two carbon atoms with only Cl and F atom

HCFC - Any saturated organic compound that contain one or two carbon atoms with only one hydrogen atom and others are Cl and F atoms

HFC - Any saturated organic compound that contain one or two carbon atoms with only one hydrogen atom and others are F atoms.

10 (b) (v): 12 marks

(vi) Select **one** class of compounds from the two classes you stated in (v) that contributes to the catalytic degradation of ozone in the upper atmosphere.

CFC or HCFC (must be selected from (v) to get marks)

10 (b) (vi): 03 marks

(vii) The slow down of industrial activities due to the Covid-19 pandemic temporarily eased the global environmental issues in many countries. Justify this statement by using **two** main global environmental issues you have learnt.

Reduction of Global warming (01): Due to the **reduction of emission of CO_2 (01)** because of **reduction of fossil fuel burning (02)** due to **limitation of industrial activities (01) and transportation (01)**.

Reduction of acid rain (01) : Reduction of emission of SO_2 (01) into the atmosphere due to decrease of **burning of coal and diesel (01+01)** for power **generation and transportation (01+01)** respectively .

or

Reduction of acid rain (01) Reduction of emission of **NO_2/NO into the atmosphere (01)** due to **decrease of fuel burning (01)** in **internal combustion engine (01)** of vehicles caused by limitation of **transportation (02)**.

Reduction of Photochemical smog (01). Reduction of emission **of NO** and **volatile hydrocarbons (01+01)** into the atmosphere from internal combustion engines/vehicles **(01)** due to limitation of transportation **(02)**.

Any two (06 x 2 = 12 marks)

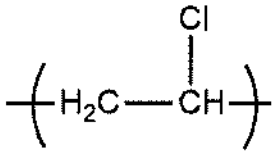
10 (b) (vii): 12 marks

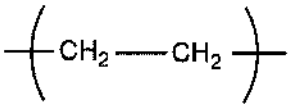
10(b): 50 marks

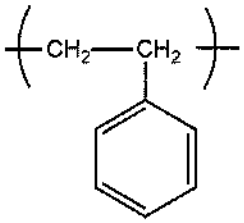
(c) The following questions are based on the polymers given below.

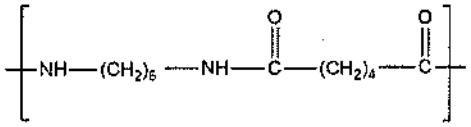
Polyvinyl chloride (PVC), Polyethylene (PE), Polystyrene (PS), Bakelite, Nylon 6.6, Polyethylene terephthalate (PET), Gutta percha

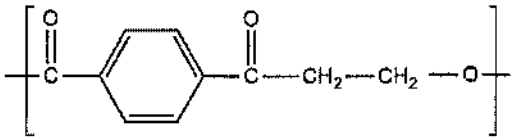
(i) Draw the repeating units of fo f the above polymers.

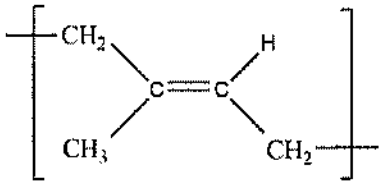
PVC 

PE 

PS 

Nylon 6,6 

PET 

Gutta percha 

Brackets are not required for award of marks.

Any four **(02 x 4 = 08)**

10 (c) (i): 08 marks

(ii) Categorize each of the above seven (7) polymers as either,

I. natural or synthetic polymers.

II. addition or condensation polymers.

	I - natural/synthetic	II - addition/condensation
PVC	synthetic	addition
PE	synthetic	addition
PS	synthetic	addition
Bakelite	synthetic	condensation
Nylon 6,6	synthetic	condensation
PET	synthetic	condensation
Gutta percha	natural	addition

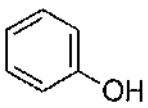
For I – Any 6 **(02 x 6 = 12)**

For II – Any 6 **(02 x 6 = 12)**

10 (c) (ii): 24 marks

(iii) Name the **two** monomers used in the formation of bakelite.

phenol and formaldehyde

OR  and HCHO

(02 x 2 = 04)

10 (c) (iii): 04 marks

(iv) Polymers can be grouped into two categories based on their thermal properties. State these **two** categories. Write to which of these categories PVC and bakelite belong.

Thermoset polymers **(02)**

Thermoplastic polymers **(02)**

Bakelite – thermoset polymer **(02)**

PVC – thermoplastic polymer **(02)**

10 (c) (iv): 08 marks

(v) Give **one** use each for **three** of the polymers given in the above list.

PVC	pipes to supply water, seat cover, electric wire covers
PE	food wrapping, garbage bags
PS	stylofoam cups, rigiform, insulating materials, packing materials
Bakelite	heat resistant parts for electric utensils, insulating materials
Nylon 6,6	clothes, fishing nets & lines, tyre threads
PET	bottles
Gutta percha	insulation, permanent tooth fillings, golf balls

Any three **(02 x 3 = 06)**

10(c) (v): 06 marks

10(c): 50 marks